



balance nuclear reaction equations

aluminum-26 undergoes positron emission
potassium-40 undergoes beta decay (electron)
uranium-238 undergoes alpha decay

electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- Sulfur hexafluoride SF_6
- Lithium nitride Li_3N
- Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- Tin (II) chloride SnCl_2
- Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- Mercury (I) chloride Hg_2Cl_2
- Potassium bromate KBrO_3
- Hydrosulfuric acid $\text{H}_2\text{S}(\text{aq})$
- Chloric acid HClO_3
- Sulfurous acid H_2SO_3

Name each of the following compounds.

- CuSO_4 Copper (II) sulfate
- AlF_3 aluminum fluoride
- HI hydrogen iodide
- NO nitrogen monoxide
- H_2Se hydroseonic acid
- HNO_3 nitric acid
- HNO_2 nitrous acid
- NaHSO_4 sodium bisulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives } 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{r} 63 \times .692 = 43.596 \rightarrow 43 \\ 65 \times .308 = 20.02 \rightarrow 20 \\ \hline 63.62 \rightarrow 64 \end{array}$$

Answers Study Guide Chemistry

NES Chemistry Exam Prep Team

Answers Study Guide Chemistry:

Study Guide for Chemical Principles [by] Steven S. Zumdahl Steven S. Zumdahl, Paul B. Kelter, 1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text. The new edition of the Study Guide includes many new worked out examples. *Organic Chemistry Study Guide with Solutions Manual* Neil E. Schore, 2007 The guide includes chapter introductions that highlight new material, chapter outlines, detailed comments for each chapter section, a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer. *Study Guide for Organic Chemistry* Seyhan N. Ege, 1999 **NES Chemistry Study Guide** NES Chemistry Exam Prep Team, 2018-02-12 Imagine a study guide actually designed for teachers. Because we know you've got a busy life we've developed a study guide that isn't like other certification materials out there. With our NEW NES Chemistry Study Guide Test Prep and Practice Questions for the National Evaluation Series Chemistry Exam you get a swift but full review of everything tested on your certification exam. FREE online resources are also included with your study guide. Imagine having FREE practice questions, digital flash cards, study cheat sheets and 35 test tips available anytime anywhere on your cell phone or tablet. Our resources will give you the push you need to pass your test the first time. Because NES Chemistry Study Guide tells you everything you need to know and nothing you don't, you'll have more time to concentrate on what's important in your life. Furthermore, NES Chemistry Study Guide gives you the opportunity to test your knowledge and assess your skills with practice tests and answers based on the actual exam. Because teaching is more than a test we give you JUST what you need to succeed. NES Chemistry Study Guide Test Prep and Practice Questions for the National Evaluation Series Chemistry Exam will provide you with a detailed overview of the NES Chemistry exam so you know exactly what to expect on test day. We'll take you through all the concepts covered on the test and give you the opportunity to test your knowledge with practice questions. Even if it's been a while since you last took a major test, don't worry we'll make sure you're more than ready. Cirrus Test Prep's NES Chemistry Study Guide Test Prep and Practice Questions for the National Evaluation Series Chemistry Exam includes a comprehensive REVIEW of Basic Principles of Matter, Atomic and Nuclear Structure, Bonding, Naming Compounds, Chemical Reactions, Thermodynamics, Solutions and Acid Base Chemistry, Scientific Inquiry and Procedures, as well as a FULL NES Chemistry practice test. About Cirrus Test Prep Developed by experienced current and former educators, Cirrus Test Prep's study materials help future educators gain the skills and knowledge needed to successfully pass their state level teacher certification exams and enter the classroom. Each Cirrus Test Prep study guide includes a detailed summary of the test's format, content and scoring, an overview of the content knowledge required to pass the exam, worked through sample questions with answers and explanations, full length practice tests including answer explanations, and unique test taking strategies with highlighted key concepts. Cirrus Test Prep's study materials ensure that new educators feel prepared on test day and beyond. [Chemical Principles Student's Study Guide & Solutions Manual](#) John Krenos, Joseph Potenza, Loretta

Jones,Lynn Koplitz,Thomas Spence,2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text **Study Guide and Full Solutions Manual** Susan McMurry,John McMurry,2002-08

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