

Chemical Reactions:

1. The following reaction $\text{PbO}_2 + \text{H}_2\text{O}_2 \rightarrow \text{PbO}_2\text{O}_3 + \text{H}_2\text{O}$ is not true. Which is not true?

A) reaction with 2.0 mol/l H_2O_2
 B) ground reaction with 3.0 mol/l H_2O_2
 C) reaction with 1.0 mol/l H_2O_2
 D) ground reaction with 2.0 mol/l H_2O_2
 E) reaction with 4.0 mol/l H_2O_2

2. A 1.0 mol/l example of the NO_x reaction is oxygenated according to the equation: $2\text{NO}_2\text{O}_2(\text{g}) \rightarrow 2\text{NO}_2\text{O}(\text{g}) + 2\text{O}_2(\text{g})$. Below, which reaction of O_2 with NO_2 will result in the same amount of NO_2 gas?

A) 0.1 mol
 B) 1.0 mol
 C) 1.2 mol
 D) 0.7 mol
 E) 2.0 mol

3. The reaction of NO_2 is step characterized by the equation: $4\text{NO}_2 + \text{O}_2 \rightleftharpoons 2\text{NO}_2\text{O}_2$. If reaction 2.0 mol/l NO_2 and 0.5 mol/l O_2 occurs, how many moles of O_2 are consumed in the reaction?

A) 0.027 mol
 B) 0.790 mol
 C) 1.000 mol
 D) 1.54 mol
 E) 2.07 mol

4. Calculated on the following reaction, v from 2000 mol/l NO_2 and 2000 mol/l O_2 is $\text{NO}_2\text{O}_2(\text{g}) + \text{NO}_2(\text{g}) \rightleftharpoons 2\text{NO}_2\text{O}(\text{g})$.

$\text{NO}_2 + \text{O}_2 \rightarrow \text{NO}_2\text{O}_2$ is not true. Which is not true?

A) reaction with 2.0 mol/l O_2
 B) ground reaction with 3.0 mol/l O_2
 C) reaction with 1.0 mol/l O_2
 D) ground reaction with 2.0 mol/l O_2
 E) reaction with 4.0 mol/l O_2

5. After the following reaction, v from 2000 mol/l NO_2 and 2000 mol/l O_2 is $\text{NO}_2\text{O}_2(\text{g}) + \text{NO}_2(\text{g}) \rightleftharpoons 2\text{NO}_2\text{O}(\text{g})$.

$\text{NO}_2 + \text{O}_2 \rightarrow \text{NO}_2\text{O}_2$ is not true. Which is not true?

A) reaction with 2.0 mol/l O_2
 B) ground reaction with 3.0 mol/l O_2
 C) reaction with 1.0 mol/l O_2
 D) ground reaction with 2.0 mol/l O_2
 E) reaction with 4.0 mol/l O_2

6. After the following reaction, v from 2000 mol/l NO_2 and 2000 mol/l O_2 is $\text{NO}_2\text{O}_2(\text{g}) + \text{NO}_2(\text{g}) \rightleftharpoons 2\text{NO}_2\text{O}(\text{g})$.

$\text{NO}_2 + \text{O}_2 \rightarrow \text{NO}_2\text{O}_2$ is not true. Which is not true?

A) reaction with 2.0 mol/l O_2
 B) ground reaction with 3.0 mol/l O_2
 C) reaction with 1.0 mol/l O_2
 D) ground reaction with 2.0 mol/l O_2
 E) reaction with 4.0 mol/l O_2

7. After the following reaction, v from 2000 mol/l NO_2 and 2000 mol/l O_2 is $\text{NO}_2\text{O}_2(\text{g}) + \text{NO}_2(\text{g}) \rightleftharpoons 2\text{NO}_2\text{O}(\text{g})$.

$\text{NO}_2 + \text{O}_2 \rightarrow \text{NO}_2\text{O}_2$ is not true. Which is not true?

A) reaction with 2.0 mol/l O_2
 B) ground reaction with 3.0 mol/l O_2
 C) reaction with 1.0 mol/l O_2
 D) ground reaction with 2.0 mol/l O_2
 E) reaction with 4.0 mol/l O_2

8. The following reaction $2\text{NO}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}_2\text{O}(\text{g})$ is not true. Which reaction of NO_2 with O_2 would the ground reaction be 2.0 mol/l NO_2 and 1.0 mol/l O_2 ?

A) 1.02 mol
 B) 0.57 mol
 C) 0.737 mol
 D) 0.19 mol
 E) 0.393 mol

9. After the following reaction, v from 2000 mol/l NO_2 and 2000 mol/l O_2 is $\text{NO}_2\text{O}_2(\text{g}) + \text{NO}_2(\text{g}) \rightleftharpoons 2\text{NO}_2\text{O}(\text{g})$.

$\text{NO}_2 + \text{O}_2 \rightarrow \text{NO}_2\text{O}_2$ is not true. Which is not true?

A) 0.099 $\times 10^{22}$ g
 B) 0.011 $\times 10^{22}$ g
 C) 4.98 $\times 10^{21}$ g
 D) 4.98 $\times 10^{22}$ g
 E) 0.179 $\times 10^{22}$ g

10. After reaction neutralization and oxidation of reaction 1000 mol/l NO_2 and 2000 mol/l O_2 is measured in treatment on the following reaction:



A) 1.07 $\times 10^{22}$ molecules
 B) 1.07 $\times 10^{22}$ molecules
 C) 4.34 $\times 10^{22}$ molecules
 D) 5.00 $\times 10^{22}$ molecules
 E) 7.17 $\times 10^{22}$ molecules

11. Calculate the mass of carbon dioxide produced from 2000 mol/l $\text{C}_2\text{H}_5\text{OH}$ in the following reaction:



A) 778 $\times 10^{-3}$ g
 B) 1.76 $\times 10^{-3}$ g
 C) 9.91 $\times 10^{-3}$ g
 D) 1.200 $\times 10^{-3}$ g
 E) 1.800 $\times 10^{-3}$ g

12. Calculate the mass of oxygen required to react with 1.0 $\times 10^{-3}$ mol of sulfur in the following reaction:



A) 1.028 $\times 10^{22}$ molecules O_2
 B) 0.258 $\times 10^{22}$ molecules O_2
 C) 1.71 $\times 10^{22}$ molecules O_2
 D) 2.402 $\times 10^{22}$ molecules O_2
 E) 0.123 $\times 10^{22}$ molecules O_2

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**James S. Mattson, Harry B.
Mark, Hubert C. MacDonald**

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