

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of a solution and the pH of a solution based on concentration.
2. Calculate pH and concentration of acids and bases and the pH of a solution based on concentration.
3. Calculate the pH of a solution (pH calculation).

## Vocabulary

- |                   |   |                  |
|-------------------|---|------------------|
| 1. autoionization | 2. hydronium ion ( $\text{H}_3\text{O}^+$ ) | 3. acid solution |
| 4. autoionization | 5. acid rain                                | 6. acid rain     |
| 7. autoionization | 8. acid rain                                | 9. pH            |

## Key Equation

- |                                    |  |
|------------------------------------|--|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ |

## Part A Completion

Write the equation for the autoionization of water. What is the concentration of  $\text{H}^+$  in pure water? What is the concentration of  $\text{OH}^-$  in pure water?

What is the pH of a solution with a concentration of  $1.0 \times 10^{-7} \text{ M}$   $\text{H}^+$ ?

What is the pH of a solution with a concentration of  $1.0 \times 10^{-7} \text{ M}$   $\text{OH}^-$ ?

What is the pH of a solution with a concentration of  $1.0 \times 10^{-7} \text{ M}$   $\text{H}^+$ ?

What is the pH of a solution with a concentration of  $1.0 \times 10^{-7} \text{ M}$   $\text{OH}^-$ ?

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What is the pH of a solution with a concentration of  $1.0 \times 10^{-7} \text{ M}$   $\text{H}^+$ ?

What is the pH of a solution with a concentration of  $1.0 \times 10^{-7} \text{ M}$   $\text{OH}^-$ ?

1.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$

2.  $1.0 \times 10^{-7} \text{ M}$

3.  $1.0 \times 10^{-7} \text{ M}$

4.  $1.0 \times 10^{-7} \text{ M}$

5.  $1.0 \times 10^{-7} \text{ M}$

6.  $1.0 \times 10^{-7} \text{ M}$

7.  $1.0 \times 10^{-7} \text{ M}$

8.  $1.0 \times 10^{-7} \text{ M}$

9.  $1.0 \times 10^{-7} \text{ M}$

10.  $1.0 \times 10^{-7} \text{ M}$

11.  $1.0 \times 10^{-7} \text{ M}$

12.  $1.0 \times 10^{-7} \text{ M}$

# 19 2 Hydrogen Ions And Acidity Work Sheet Answers



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