



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Define acid and base and relate them to the pH of a solution.
- Define strong and weak acids.

Key Vocabulary

- | | |
|---|---|
| <ul style="list-style-type: none">hydrogen ionacidbasestrong acidweak acidpH | <ul style="list-style-type: none">hydrogen ion concentrationacidicbasicneutraltitrationtitration curve |
|---|---|

Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

Key Diagrams

- Titration curve
- Titration curve

Points of Consideration

Acidic and basic properties are determined by the concentration of hydrogen ions in a solution. The concentration of hydrogen ions is measured in moles per liter.

The concentration of hydrogen ions in a solution is determined by the concentration of other ions in the solution. The concentration of other ions in the solution is determined by the concentration of other ions in the solution.

The concentration of hydrogen ions in a solution is determined by the concentration of other ions in the solution. The concentration of other ions in the solution is determined by the concentration of other ions in the solution.

The concentration of hydrogen ions in a solution is determined by the concentration of other ions in the solution. The concentration of other ions in the solution is determined by the concentration of other ions in the solution.

Key Equations

- | | |
|--------------------------------------|--------------------------------------|
| $\text{pH} = -\log_{10}[\text{H}^+]$ | $\text{pH} = \log_{10}[\text{OH}^-]$ |
| $\text{pH} = 7$ | $\text{pH} = 14$ |
| $\text{pH} = 1$ | $\text{pH} = 13$ |
| $\text{pH} = 2$ | $\text{pH} = 12$ |
| $\text{pH} = 3$ | $\text{pH} = 11$ |
| $\text{pH} = 4$ | $\text{pH} = 10$ |
| $\text{pH} = 5$ | $\text{pH} = 9$ |
| $\text{pH} = 6$ | $\text{pH} = 8$ |
| $\text{pH} = 7$ | $\text{pH} = 7$ |
| $\text{pH} = 8$ | $\text{pH} = 6$ |
| $\text{pH} = 9$ | $\text{pH} = 5$ |
| $\text{pH} = 10$ | $\text{pH} = 4$ |
| $\text{pH} = 11$ | $\text{pH} = 3$ |
| $\text{pH} = 12$ | $\text{pH} = 2$ |
| $\text{pH} = 13$ | $\text{pH} = 1$ |
| $\text{pH} = 14$ | $\text{pH} = 0$ |

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