

Name Key

1. Hydrogen,  $H_2$ , is used as rocket fuel. The hydrogen is burned in oxygen,  $O_2$ , according to the following equation. What is the enthalpy change if 200.0 g of hydrogen,  $H_2$ , is burned?



$$\frac{200.0 \text{ g } H_2}{2.016 \text{ g } H_2} \times \frac{1 \text{ mol } H_2}{1} = 99.2 \text{ mol } H_2$$

$$99.2 \text{ mol } H_2 \times \frac{-484 \text{ kJ}}{2 \text{ mol } H_2} = -24,006 \text{ kJ}$$

2. Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide.



What is  $\Delta H$  for the following reaction?



3. Following is the equation for the combustion of acetone:



How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

$$\frac{-853.6 \text{ kcal}}{2 \text{ mol } C_3H_6O} \times 4.5 \text{ mol } C_3H_6O = -1920.6 \text{ kcal}$$

4. The specific heat of iron is  $0.110 \text{ cal} / \text{g} \cdot ^\circ\text{C}$ . How much heat must be added to 5.2 g of iron at  $25^\circ\text{C}$  to raise its temperature to  $225^\circ\text{C}$ ?

$$\text{heat} = q = m c \Delta T = (5.2 \text{ g}) \left( \frac{0.110 \text{ cal}}{\text{g} \cdot ^\circ\text{C}} \right) (225 - 25^\circ\text{C}) = 114.4 \text{ cal}$$

5. Hydrazine,  $N_2H_4$ , is a colorless liquid used as a rocket fuel. What is the enthalpy change for the process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate  $\Delta H$  for reaction (1) given the data below.



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Solid-State Chemistry Ashok Kumar Jha, 2023-08-04 Solid state reactions have been growing in importance in the preparation of solids, crystal growth, and single crystals of elements. In this new volume, the author discusses in depth recent advances in solid state chemistry, the latest techniques of characterization of solids, and several new dimensions of solid state problems. Solid state reactions, their kinetics, phase transitions, and magnetic properties are discussed in detail, along with electrical properties, semiconductors, metals, and insulators, with examples and explanations. Dielectric, piezoelectric, and ferroelectric properties of solids have been considered as well. Different theories of catalysis and adsorption isotherms are covered, and the author includes several experimental findings as well. And, of course, preparation of new solids by taking advantage of recent experimental techniques have also been included in this volume. This volume provides a complete description of solid state reactions, electrical conductivity, characterization of solids, organic solids, and intercalation between layers of solids. The band theories of metals, semiconductors, and insulators are covered. In characterization of solids, TGA, DTA, DSC, FTIR, and PXRD techniques have been described. Preparation and properties of organic solids are also explained. Dielectric properties, photovoltaic effects, piezoelectricity, ferroelectricity, and antiferroelectricity have been dealt with, giving relatable examples. A number of excitation energy levels giving rise to luminescence and photoluminescence have been explained under the topic of optical processes. Diamagnetic, paramagnetic, and ferromagnetic behaviors of solids are also explained.

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