

CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY**Solutions****Section 15.1 What are solutions?**

In your textbook, read about the characteristic of solutions.

Underline the terms below just now to complete the passage.

solvent	liquid	solvent
insoluble	solid	solvent

After reading (1) _____ solution _____ of solute is dissolved in solvent.

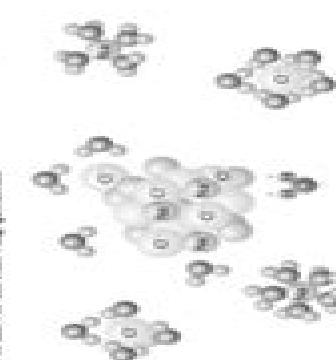
(2) _____ solute _____ is dissolved in the (3) _____ solvent _____.

(4) _____ solubility _____ is the ability of a substance to dissolve in another substance. A substance that does not dissolve is (5) _____ insoluble _____.

(6) _____ liquid _____ solubility of the most common type of solvents. If one liquid is soluble in another liquid, such as water and oil, water is (7) _____ miscible _____ in oil. However, if one liquid is insoluble in another, the liquids are (8) _____ immiscible _____.

Read about solubility in aqueous solution in your textbook.

Use the diagram to answer the following questions.



The diagram shows the hydration of solid sodium chloride to form an aqueous solution.

Use the diagram to answer the following questions.

1. In what is solute in which the solute is a solid? What is solute?

Solute is the process of surrounding solute particles with solvent particles to form a solution.

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Density, Matter and Change • Chapter 11

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Section 15.2 **STUDY GUIDE FOR CONTENT MASTERY****Section 15.1** **Answers**

16. An sodium chloride dissolved in water. What happens to the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.

17. Unlike the attraction of the water molecules around the sodium ions and chloride ions.

Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atom). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solution process to occur.

19. In a glass of water, the rate of solution may be increased by stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent.

In your textbook, read about heat of solution, solubility and factors that affect solubility for each solute and solvent, with their graphs.

True 14. The overall energy change that occurs when a solute dissolves in a solvent is heat of solution.

False 15. Solubility is a property of the solute, amount of which the solution is a given amount of solvent at a specified temperature and pressure.

False 16. Solvent molecules in liquid as the solute molecules in liquid are the only molecules that move.

True 17. In a saturated solution, solute and solvate molecules are in equilibrium.

True 18. Additional solute can be dissolved in an unsaturated solution.

True 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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Chapter 12

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