

## Chemistry 2<sup>nd</sup> Semester Final Review KEY

\*This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.\*

Bring your book to school by Monday for turn in (seniors should bring book to bookstore by themselves ASAP)

Gas Laws (Ch. 13):

Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)

Practice questions:

1. If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)

$$V_1 = 3.8 \text{ L}; P_1 = 2.71 \text{ atm}$$

$$V_2 = 1.47 \text{ L}; P_2 = ? \text{ atm}$$

Since all units are standard (L, atm), we do not need to convert the units before proceeding.

$$(P_1 V_1) / T_1 = (P_2 V_2) / T_2$$

T goes away (not mentioned in problem)

$$P_1 V_1 = P_2 V_2$$

Isolate for P2:  $P_2 = (P_1 V_1) / V_2$

$$P_2 = (2.71 \text{ atm} \times 3.8 \text{ L}) / (1.47 \text{ L}) = 7.0 \text{ atm} \text{ (L cancelled and it's 2 sig figs)}$$

2. A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)

$$V_1 = 1.49 \text{ L}, T_1 = 34.75 \text{ degrees Celsius}$$

$$V_2 = ? \text{ L}, T_2 = 78.41 \text{ degrees Celsius}$$

We need to convert from Celsius to Kelvins before proceeding.

$$T_1 = 34.75 + 273 = 307.75 \text{ K}$$

$$T_2 = 78.41 + 273 = 351.41 \text{ K}$$

$$(P_1 V_1) / T_1 = (P_2 V_2) / T_2$$

Get rid of P

$$V_1 / T_1 = V_2 / T_2$$

Isolate V2:  $V_2 = (V_1 / T_1) T_2$

$$V_2 = (1.49 \text{ L} / 307.75 \text{ K}) (351.41 \text{ K}) = 1.70 \text{ L}$$

3. What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?

The combined gas law won't work here. You only have one condition and nothing changes.

$PV = nRT$  will work. ( $n$ = moles and  $R = 0.08026 \text{ L atm/K mol}$ )

$P = 1.04 \text{ atm}$ ,  $V = ? \text{ L}$ ,  $n =$  OOPS we need to convert from grams to moles!,  $T = 84.7 \text{ degrees Celsius} + 273 = 357.7 \text{ K}$

$n$ : Hydrogen gas is diatomic!

$$8.47 \text{ g H}_2 \times (1 \text{ mol H}_2 / 2.02 \text{ g H}_2) = 4.193069307 \text{ mol H}_2$$

$$\text{Isolate V: } V = (nRT) / P = (4.193069307 \text{ mol H}_2 \times 0.08026 \text{ L atm/K mol} \times 357.7 \text{ K}) / 1.04 \text{ atm} = 118 \text{ L}$$

4. What volume is occupied by 56.75 g of oxygen gas at STP?

# 2nd Semester Chemistry Final Exam Study Guide

**James E. Brady**

## 2nd Semester Chemistry Final Exam Study Guide:

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Brady,1975-04 *United States Air Force Academy* United States Air Force Academy, *The 2004 Guide to the Evaluation of Educational Experiences in the Armed Services* American Council on Education Staff,2004-10-27 **EC Study Guide to Environment-related Courses** Institute for European Environmental Policy,Commission of the European Communities.

Directorate-General Environment, Nuclear Safety and Civil Protection,1993 El-Hi Textbooks in Print ,1977 Annual Catalogue United States Air Force Academy,1984 *Catalog* James Millikin University,1912 *The National Guide to Educational Credit for Training Programs* American Council on Education,2005 Highlights over 6 000 educational programs offered by business labor unions schools training suppliers professional and voluntary associations and government agencies

*Nursing School Entrance Exam Review* William H. Brown, Frederick A. Bettelheim, Jerry March,2004-04 **Proceedings** Bassam Z. Shakhshiri,1973 **How to Pass High on the U.S. Coast Guard Academy Tests** Arco Publishing

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## **Table of Contents 2nd Semester Chemistry Final Exam Study Guide**

1. Understanding the eBook 2nd Semester Chemistry Final Exam Study Guide
  - The Rise of Digital Reading 2nd Semester Chemistry Final Exam Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying 2nd Semester Chemistry Final Exam Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 2nd Semester Chemistry Final Exam Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from 2nd Semester Chemistry Final Exam Study Guide

- Personalized Recommendations
- 2nd Semester Chemistry Final Exam Study Guide User Reviews and Ratings
- 2nd Semester Chemistry Final Exam Study Guide and Bestseller Lists

5. Accessing 2nd Semester Chemistry Final Exam Study Guide Free and Paid eBooks

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- 2nd Semester Chemistry Final Exam Study Guide eBook Subscription Services
- 2nd Semester Chemistry Final Exam Study Guide Budget-Friendly Options

6. Navigating 2nd Semester Chemistry Final Exam Study Guide eBook Formats

- ePUB, PDF, MOBI, and More
- 2nd Semester Chemistry Final Exam Study Guide Compatibility with Devices
- 2nd Semester Chemistry Final Exam Study Guide Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of 2nd Semester Chemistry Final Exam Study Guide
- Highlighting and Note-Taking 2nd Semester Chemistry Final Exam Study Guide
- Interactive Elements 2nd Semester Chemistry Final Exam Study Guide

8. Staying Engaged with 2nd Semester Chemistry Final Exam Study Guide

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers 2nd Semester Chemistry Final Exam Study Guide

9. Balancing eBooks and Physical Books 2nd Semester Chemistry Final Exam Study Guide

- Benefits of a Digital Library
- Creating a Diverse Reading Collection 2nd Semester Chemistry Final Exam Study Guide

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine 2nd Semester Chemistry Final Exam Study Guide

- Setting Reading Goals 2nd Semester Chemistry Final Exam Study Guide
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of 2nd Semester Chemistry Final Exam Study Guide

- Fact-Checking eBook Content of 2nd Semester Chemistry Final Exam Study Guide
- Distinguishing Credible Sources

### 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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