

1. CLASSIFICATION OF MATTER

- Matter can exist in three physical states: solid, liquid and gas.



- Various properties of matter are used to differentiate between the two.
- Matter is classified into pure substance and mixture.
- Pure substance is further classified into element and compound.
- Mixture is further classified into homogeneous mixture and heterogeneous mixture.
- Homogeneous mixture is further classified into solution and alloy.
- Heterogeneous mixture is further classified into suspension and colloid.
- Colloid is a mixture of two or more substances in which the particles are not visible to the naked eye.
- Suspension is a mixture of two or more substances in which the particles are visible to the naked eye.

2. PROPERTIES OF MATTER AND THEIR MEASUREMENT

- Properties of matter are physical and chemical properties.
- Physical properties are those which can be measured without changing the identity of the substance.
- Chemical properties are those which can be measured by changing the identity of the substance.
- Colour, odour, melting point, density, etc. are some of the physical properties.
- Combustion, corrosion, etc. are some of the chemical properties.
- Quantitative measurement of physical properties is necessary for scientific study.
- Measurement of mass, length, time, etc. are some of the physical properties.
- Measurement of temperature, etc. are some of the chemical properties.

UNIT CONVERSIONS AND DIMENSIONS

10 ³ grams	10 ³ grams	10 ³ grams	10 ³ grams	10 ³ grams
10 ³ kg	10 ³ kg	10 ³ kg	10 ³ kg	10 ³ kg

UNIT CONVERSIONS

- 1 kg = 1000 g, 1 m = 1000 mm, 1 L = 1000 mL, 1 m = 1000 mm, 1 m = 1000 mm

3. SIGNIFICANT FIGURES

Significant figures are those digits in a number which are known with certainty, plus one additional uncertain or doubtful digit.

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Rules for determining the number of significant figures:

- Non-zero digits are significant.
- Zeros between non-zero digits are significant.
- Zeros at the end or right of a number are significant provided they are on the right side of a decimal point.
- Leading zeros are not significant.
- In a number written in scientific notation, all digits are significant.
- Exact values are the numbers of atoms, molecules, etc. for a substance.
- Rounding off the significant figures of a number to the least number of significant figures.
- In addition and subtraction of significant figures, the result should have the same number of significant figures as the least number of significant figures in the original numbers.
- In multiplication and division of significant figures, the result should have the same number of significant figures as the least number of significant figures in the original numbers.

Aakash Chemistry Lab Manual

Helen Mango



Aakash Chemistry Lab Manual:

Quantitative Analysis Laboratory Manual Dr. Bupinderpal Singh ,Dr. Bikramjit Singh,Mr. Akash Bans,2025-06-29

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M. van Koppen,2010 *Practical/Laboratory Manual Chemistry Class - XI* Er. Meera Goyal,2021-05-29 1 Basic Laboratory Techniques 1 To cut a glass tube or glass rod 2 To bend the glass rod at an angle 3 To draw a glass jet from a glass tube 4 To bore a cork and fit a glass tube into it Viva Voce 2 Characterisation and Purification of Chemical Substances 1 To determine the melting point of the given unknown organic compound and its identification simple laboratory technique Viva Voce 2 To determine the boiling point of a given liquid when available in small quantity simple laboratory method Viva Voce 3 To prepare crystals of pure potash alum $K_2SO_4 \cdot Al_2(SO_4)_3 \cdot 24H_2O$ from the given impure sample 4 To prepare the pure crystals of copper sulphate from the given crude sample 5 To prepare pure crystals of benzoic acid from a given impure sample Viva Voce 3 Measurement of pH Values 1 To determine the pH value of vegetable juices fruit juices tap water and washing soda by using universal pH paper 2 To determine and compare the pH values of solutions of strong acid HCl and weak acid CH_3COOH of same concentration 3 To study the pH change in the titration of strong base Vs strong acid by using universal indicator paper 4 To study the pH change by common ion CH_3COO^- ion in case of weak acid CH_3COOH 5 To determine the change in pH value of weak base NH_4OH in presence of a common ion NH_4^+ Viva Voce 4 Chemical Equilibrium 1 To study the shift in equilibrium between ferric ions and thiocyanate ions by changing the concentrations of either of the ions 2 To study the shift in equilibrium between Co^{2+} and Cl^- ions by changing the concentrations of either of the ions Viva Voce 5

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