

AP CHEMISTRY

Volume 1: Your Second Year Study

Properties of Solids

- **Crystalline Solids** have a regular, repeating arrangement of particles. They have sharp melting points and distinct X-ray diffraction patterns.
- **Amorphous Solids** lack a long-range order. They soften over a range of temperatures and do not have sharp X-ray diffraction patterns.
- **Unit Cells** are the smallest repeating units of a crystal lattice. Common types include simple cubic, body-centered cubic (BCC), and face-centered cubic (FCC).
- **Packing Efficiency** is the fraction of space occupied by atoms in a unit cell. FCC and HCP have the highest efficiency at approximately 74%.
- **Intermolecular Forces** (IMFs) determine the strength of the solid. Stronger IMFs lead to higher melting points and greater hardness.
- **Network Solids** consist of atoms bonded in a continuous 3D network (e.g., diamond, silicon dioxide). They are extremely hard and have very high melting points.
- **Hydrogen Bonding** is a special type of IMF that occurs between molecules containing N-H, O-H, or F-H bonds. It significantly affects the properties of water and other hydrogen-bonded liquids.

Types of Chemical Calculations

Key to Answer on the Exam

- **Stoichiometry**: Calculations involving the relative amounts of reactants and products in a chemical reaction.
- **Thermochemistry**: Calculations involving heat energy changes during chemical reactions.
- **Equilibrium**: Calculations involving the concentrations of reactants and products at equilibrium.
- **Acid-Base**: Calculations involving pH, pKa, and the concentrations of acids and bases.
- **Solubility**: Calculations involving the concentrations of ions in saturated solutions.
- **Electrochemistry**: Calculations involving the relationship between electrical energy and chemical reactions.

Periodic Table & Trends

The periodic table is color-coded to show trends. Electronegativity increases from bottom-left to top-right. Ionization energy follows a similar trend. Atomic radius increases from top-left to bottom-right.

- **Electronegativity** is the ability of an atom to attract electrons in a bond. It increases across a period and decreases down a group.
- **Ionization Energy** is the energy required to remove an electron from a gaseous atom. It increases across a period and decreases down a group.
- **Atomic Radius** is the distance from the nucleus to the outer edge of the electron cloud. It increases down a group and across a period.
- **Electron Affinity** is the energy change when an electron is added to a gaseous atom. It generally increases across a period.
- **Effective Nuclear Charge (Z_{eff})** is the net positive charge experienced by an electron in a multi-electron atom.

Electronegativity Table

Electronegativity (Pauling Scale)

1. Non-metals have high electronegativity values (3.0 or greater).

2. Metals have low electronegativity values (1.0 or less).

3. Electronegativity increases from left to right across a period.

4. Electronegativity decreases from top to bottom down a group.

5. Fluorine is the most electronegative element (4.0).

6. Francium is the least electronegative element (0.7).

7. Noble gases have no electronegativity values listed.

8. Electronegativity is a relative measure of an atom's ability to attract electrons.

9. Electronegativity is used to predict the polarity of a bond.

10. Electronegativity is used to predict the type of chemical bond that will form between two atoms.

11. Electronegativity is used to predict the strength of a chemical bond.

12. Electronegativity is used to predict the reactivity of an element.

13. Electronegativity is used to predict the acidity of a compound.

14. Electronegativity is used to predict the basicity of a compound.

15. Electronegativity is used to predict the solubility of a compound.

16. Electronegativity is used to predict the boiling point of a compound.

17. Electronegativity is used to predict the melting point of a compound.

18. Electronegativity is used to predict the density of a compound.

19. Electronegativity is used to predict the color of a compound.

20. Electronegativity is used to predict the odor of a compound.

21. Electronegativity is used to predict the taste of a compound.

22. Electronegativity is used to predict the toxicity of a compound.

23. Electronegativity is used to predict the flammability of a compound.

24. Electronegativity is used to predict the explosiveness of a compound.

25. Electronegativity is used to predict the stability of a compound.

26. Electronegativity is used to predict the shelf life of a compound.

27. Electronegativity is used to predict the storage conditions of a compound.

28. Electronegativity is used to predict the handling of a compound.

29. Electronegativity is used to predict the disposal of a compound.

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