

Chemical Reactions:

1. The following reaction $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ is given. What is the ΔH°_f for the reaction?

A) greater than 2.00 kcal/mol
 B) greater than 1.00 kcal/mol
 C) less than 1.00 kcal/mol
 D) less than 2.00 kcal/mol
 E) greater than 4.00 kcal/mol

2. A 1.0-molar example of H_2CO_3 reacts with excess CaCO_3 according to the equation:

$$2\text{H}_2\text{CO}_3(\text{aq}) \rightarrow 2\text{H}_2\text{O} + \text{CO}_2(\text{g})$$
 Below, which molalities of CO_2 are possible when 1.00 mol of H_2CO_3 reacts?

A) 0.0 mol
 B) 1.0 mol
 C) 1.2 mol
 D) 1.7 mol
 E) 2.0 mol

3. The reaction of NO_2 is step characterized by the equation: $4\text{NO}_2 \rightarrow 3\text{O}_2 + 2\text{N}_2\text{O}_5$. If 0.01 mol of NO_2 reacts with 1.00-molar examples of O_2 , how many moles of O_2 are consumed when 0.001 mol of NO_2 reacts?

A) 0.001 mol
 B) 0.002 mol
 C) 0.003 mol
 D) 0.004 mol
 E) 0.005 mol

4. Calculated are the following reactions, or from the experiments, are unknowns, calculated:
 $\text{H}_2\text{O}_2 + \text{H}_2\text{O}_2 \rightarrow \text{H}_2\text{O}_2(\text{aq}) + 2\text{H}_2\text{O}(\text{g})$

What is ΔH°_f for the reaction of 1.00 mol of H_2O_2 with 1.00 mol of H_2O_2 ?

A) 0°C
 B) 25°C
 C) 50°C
 D) 75°C
 E) 100°C

5. Calculated are the following reactions, or from the experiments, are unknowns, calculated:
 $\text{H}_2\text{O}_2 + \text{H}_2\text{O}_2 \rightarrow \text{H}_2\text{O}_2(\text{aq}) + 2\text{H}_2\text{O}(\text{g})$

What is ΔS°_f for the reaction of 1.00 mol of H_2O_2 with 1.00 mol of H_2O_2 ?

A) -1.00×10^{22} $\text{mol}^{-1}\text{K}^{-1}$
 B) -1.00×10^{23} $\text{mol}^{-1}\text{K}^{-1}$
 C) -1.00×10^{24} $\text{mol}^{-1}\text{K}^{-1}$
 D) -1.00×10^{25} $\text{mol}^{-1}\text{K}^{-1}$
 E) -1.00×10^{26} $\text{mol}^{-1}\text{K}^{-1}$

6. The following reaction $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ is given. Below, which number of unknowns would the product exceed when 1.00 mol of hydrogen and oxygen react?

A) 1.00 mol
 B) 1.27 mol
 C) 1.737 mol
 D) 2.00 mol
 E) 2.943 mol

7. Calculate the following reaction enthalpy of reaction:
 $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CH}_3\text{OH} + \text{H}_2\text{O}$
 What mass of hydrogen (H_2) is required to react completely with 1.00 g of C_2H_6 ?

A) 0.001×10^{22} g
 B) 0.01×10^{22} g
 C) 0.001×10^{23} g
 D) 0.01×10^{23} g
 E) 0.100×10^{23} g

8. Below, reaction mechanisms and equations of reaction are all the elements of C_2H_6 with propane as reactant are the following reactions:
 $\text{C}_2\text{H}_6(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{g}) + \text{CH}_3\text{OH}(\text{g})$

A) 1.00×10^{22} $\text{mol}^{-1}\text{K}^{-1}$
 B) 1.00×10^{23} $\text{mol}^{-1}\text{K}^{-1}$
 C) 1.00×10^{24} $\text{mol}^{-1}\text{K}^{-1}$
 D) 1.00×10^{25} $\text{mol}^{-1}\text{K}^{-1}$
 E) 1.00×10^{26} $\text{mol}^{-1}\text{K}^{-1}$

9. Calculate the mass of carbon dioxide produced from 1.00 g of octane, C_8H_{18} , in the following reaction:



10. Calculate the molarities of oxygen required to react with 1.00 g of sulfur in the following reaction:



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Clemens Wendtner

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