

## Chemistry 2<sup>nd</sup> Semester Final Review KEY

\*This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.\*

Bring your book to school by Monday for turn in (seniors should bring book to bookstore by themselves ASAP)

Gas Laws (Ch. 13):

Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)

Practice questions:

1. If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)

$$V_1 = 3.8 \text{ L}; P_1 = 2.71 \text{ atm}$$

$$V_2 = 1.47 \text{ L}; P_2 = ? \text{ atm}$$

Since all units are standard (L, atm), we do not need to convert the units before proceeding.

$$(P_1 V_1) / T_1 = (P_2 V_2) / T_2$$

T goes away (not mentioned in problem)

$$P_1 V_1 = P_2 V_2$$

Isolate for P2:  $P_2 = (P_1 V_1) / V_2$

$$P_2 = (2.71 \text{ atm} \times 3.8 \text{ L}) / (1.47 \text{ L}) = 7.0 \text{ atm} \text{ (L cancelled and it's 2 sig figs)}$$

2. A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)

$$V_1 = 1.49 \text{ L}, T_1 = 34.75 \text{ degrees Celsius}$$

$$V_2 = ? \text{ L}, T_2 = 78.41 \text{ degrees Celsius}$$

We need to convert from Celsius to Kelvins before proceeding.

$$T_1 = 34.75 + 273 = 307.75 \text{ K}$$

$$T_2 = 78.41 + 273 = 351.41 \text{ K}$$

$$(P_1 V_1) / T_1 = (P_2 V_2) / T_2$$

Get rid of P

$$V_1 / T_1 = V_2 / T_2$$

Isolate V2:  $V_2 = (V_1 / T_1) T_2$

$$V_2 = (1.49 \text{ L} / 307.75 \text{ K}) (351.41 \text{ K}) = 1.70 \text{ L}$$

3. What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?

The combined gas law won't work here. You only have one condition and nothing changes.

$PV = nRT$  will work. ( $n$  = moles and  $R = 0.08026 \text{ L atm/K mol}$ )

$P = 1.04 \text{ atm}$ ,  $V = ? \text{ L}$ ,  $n =$  OOPS we need to convert from grams to moles!,  $T = 84.7 \text{ degrees Celsius} + 273 = 357.7 \text{ K}$

$n$ : Hydrogen gas is diatomic!

$$8.47 \text{ g H}_2 \times (1 \text{ mol H}_2 / 2.02 \text{ g H}_2) = 4.193069307 \text{ mol H}_2$$

$$\text{Isolate V: } V = (nRT) / P = (4.193069307 \text{ mol H}_2 \times 0.08026 \text{ L atm/K mol} \times 357.7 \text{ K}) / 1.04 \text{ atm} = 118 \text{ L}$$

4. What volume is occupied by 56.75 g of oxygen gas at STP?

# 2nd Semester Chemistry Final Exam Study Guide

**Brown University**

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Directory of Distance Learning Opportunities Modoc Press, Inc.,2003-02-28 This book provides an overview of current K 12 courses and programs offered in the United States as correspondence study or via such electronic delivery systems as satellite cable or the Internet The Directory includes over 6 000 courses offered by 154 institutions or distance learning consortium members Following an introduction that describes existing practices and delivery methods the Directory offers three indexes Subject Index of Courses Offered by Level Course Level Index Geographic Index All information was supplied by the institutions Entries include current contact information a description of the institution and the courses offered grade level and admission information tuition and fee information enrollment periods delivery information equipment requirements credit and grading information library services and accreditation

### **Curriculum Handbook with General Information**

**Concerning ... for the United States Air Force Academy** United States Air Force Academy,1996 **Instructor's manual and solutions to problems to accompany General chemistry: principles and structures** James E.

Brady,1975-04 *United States Air Force Academy* United States Air Force Academy, *The 2004 Guide to the Evaluation of Educational Experiences in the Armed Services* American Council on Education Staff,2004-10-27 **EC Study Guide to Environment-related Courses** Institute for European Environmental Policy,Commission of the European Communities.

Directorate-General Environment, Nuclear Safety and Civil Protection,1993 El-Hi Textbooks in Print ,1977 Annual Catalogue United States Air Force Academy,1984 *Catalog* James Millikin University,1912 *The National Guide to Educational Credit for Training Programs* American Council on Education,2005 Highlights over 6 000 educational programs offered by business labor unions schools training suppliers professional and voluntary associations and government agencies

*Nursing School Entrance Exam Review* William H. Brown, Frederick A. Bettelheim, Jerry March,2004-04 **Proceedings** Bassam Z. Shakhshiri,1973 **How to Pass High on the U.S. Coast Guard Academy Tests** Arco Publishing

Company,1959 *Catalogue of the Officers and Students of Brown University* Brown University,1911 Catalogue of the Officers and Students Brown University,1917 *Catalogue* Brown University,1911 *How to Pass Annapolis Entrance Exams* Arco Publishing Company,1959 *Planning Reports, Clippings and Other Materials Relating to Community High School and Earthworks (formerly Pioneer II)* Experimental High School Ann Arbor Public Schools, **Prep for College Adm Exam Eol 2000 G 12** Holt Rinehart & Winston,1999-05-12 High School Chemistry Teachers Magazine ,1967

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