

- Structure, Hybridization, Resonance, Aromaticity
  - # of bonds-1= $sp$  (e.g. 3 bonds =  $sp^2$ )
  - Aromaticity=  $4n+2$  electrons
  - Strained ring causes greater heat of combustion
- Acids and Bases
  - Lewis acid= accepts electron pair
  - Higher percentage of s character means more acidic hydrogen
- Stereoisomerism
  - Constitutional Isomers
    - Isomers that have different atom connectivity
  - Stereoisomers
    - Same connectivity but different arrangement in space
    - Enantiomers
      - Stereoisomers that are non-superimposable mirror reflections
    - Diastereomers
      - Stereoisomers that are not mirror reflections of each other
  - Stereocenter=  $sp^3$  carbon bonded to four different substituents
  - Assigning R/S configuration
    - If lowest group is pointing away, determine normally
    - If lowest group is not pointing away, exchange with group pointing away and determine reverse

- Nucleophilic Substitutions and Eliminations

◦ Alkyl Halide	◦ Strong Base	◦ Weak Base	◦ Poor Nucleophile
◦ 1	◦ $S_N2$	◦ $S_N2$	◦ No rxn.
◦ 2	◦ Mostly E2	◦ $S_N2$	◦ $S_N1/E1$ in polar, protic solvents
◦ 3	◦ E2	◦ $S_N1/E1$ in polar, protic solvent	◦ $S_N1/E1$ in polar, protic solvents

- Relative strength of nucleophiles
  - Negatively charged nucleophiles react faster in substitution reactions than neutral nucleophiles
  - Nucleophilic strength is in the same order as base strength when the nucleophilic atoms are in the same period of the periodic table
  - Nucleophilic strength increases with increasing atom size when the nucleophilic atoms are in the same column of the periodic table
- Leaving group ability
  - Increases with decreasing base strength
- Base strength and structure

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


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