

1. Bases react with acids to produce water and salt and this type of reaction is called a neutralization reaction.
2. What is the chemical name of H_2S ? hydrosulfuric acid ide → hydro-ic
3. What is the chemical name of H_2SO_3 ? sulfurous acid ite → ous
4. What is the chemical name of H_2SO_4 ? sulfuric acid ate → ic
5. What is the chemical formula for hydrophosphoric acid? H_3P
6. What is the chemical formula for phosphoric acid? H_3PO_4
7. What is the chemical name of $\text{Ca}(\text{OH})_2$? calcium hydroxide
8. Arrhenius Acids & bases ionize to form H^+ and OH^- respectively is an assumption of what theory?
9. Bronsted-Lowry acids donate a proton while bases accept a proton is an assumption of what theory?
10. Circle the donate Bronsted-Lowry acid and box the accept Bronsted-Lowry base in the following reaction.

$$\text{NH}_3(\text{l}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{NH}_4^+(\text{aq}) + \text{OH}^-(\text{aq})$$

base
acid
conj. acid
conj. base
11. A student places a 0.34 gram sample of sodium in an acidic solution and a violent reaction occurs producing what gas? hydrogen gas

Label the following as an Arrhenius Acid or an Arrhenius Base

12. acid HCl
13. base NaOH
14. base $\text{Mg}(\text{OH})_2$
15. acid H_2SO_4

Label the following reactions as:

- precipitation (& circle the precipitate)
 - neutralization or
16. neutralization $\text{HNO}_3 + \text{NaOH} \rightarrow \text{NaNO}_3 + \text{H}_2\text{O}$
acid base salt water
 17. neutralization $\text{HCl} + \text{KOH} \rightarrow \text{H}_2\text{O} + \text{KCl}$
 18. precipitate $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$

Acid Base Study Guide Answer Key

Paul B. Kelter, Steven S. Zumdahl



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
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