

Acids and Bases

Overview of Acids and Bases

Acids and Bases

Acids dissociate to release **proton** (H^+)

Bases dissociate to accept **proton** (H^+)



Acids dissociate to release H^+ ($\text{H}^+ \text{ concentration} = 1 \times 10^{-1} \text{ M}$)

Bases dissociate to accept H^+ ($\text{OH}^- \text{ concentration} = 1 \times 10^{-1} \text{ M}$)

What does H^+ concentration and OH^- concentration tell us?

Concentrations and Strengths

Acids are **strong acids** if H^+ concentration is $\geq 1 \times 10^{-1} \text{ M}$
Acids are **weak acids** if H^+ concentration is $< 1 \times 10^{-1} \text{ M}$



What does H^+ concentration tell us?

Acids have **lower** H^+ concentration than bases
and acids have **higher** H^+ concentration than neutral

Neutral: H^+ concentration is **intermediate** of acids and bases

Weak Acids and Weak Bases are from **Polar Properties**
What if there is no proton available?

Concentrations and Strengths

Acids with **higher** H^+ concentration
have **stronger** acids



Acids with **lower** H^+ concentration are described as **weaker acids**.

Bases with **higher** OH^- concentration are **stronger bases**

Acids with **lower** OH^- concentration are **weaker acids**

What does H^+ concentration tell us?

Acids with **higher** H^+ concentration are **stronger acids**

Acids with **lower** H^+ concentration are **weaker acids**

What does OH^- concentration tell us?

Bases with **higher** OH^- concentration are **stronger bases**

Bases with **lower** OH^- concentration are **weaker bases**

Acid Strengths

What does a lower pH mean?

Generally speaking, the lower the pH , the stronger the acid

$\text{pH} = -\log[\text{H}^+]$

Acids with **lower** pH have **higher** H^+ concentration

Acids with **higher** pH have **lower** H^+ concentration

Acids with **lower** pH are **stronger acids**

Acids with **higher** pH are **weaker acids**

What does a lower pH mean?

Acids with **lower** pH have **higher** H^+ concentration

Acids with **higher** pH have **lower** H^+ concentration

Acids with **lower** pH are **stronger acids**

Acids with **higher** pH are **weaker acids**

What does a lower pH mean?

Acids with **lower** pH have **higher** H^+ concentration

Acids with **higher** pH have **lower** H^+ concentration

Acids with **lower** pH are **stronger acids**

Acids with **higher** pH are **weaker acids**

Acid Strengths and pH

The pH value of strong acids (HCl) and strong bases (NaOH) are shown below

$$\text{HCl} = \frac{[\text{H}^+][\text{Cl}^-]}{[\text{HCl}]} = 10^{-1} \quad \text{NaOH} = \frac{[\text{OH}^-][\text{Na}^+]}{[\text{NaOH}]} = 10^{-1}$$

What does it mean to **strong** and **weak** in this context?

To **compute** the **concentration**, **logarithm** is used, which is the **pH** of the **acid** and

$$\text{pH} = -\log[\text{H}^+] = 10 \rightarrow [\text{H}^+] = 10^{-1}$$

As the **strength** of the **acid**, the **lower** the **pH** value, the **stronger** the acid

Acids are **pH** values of **lower** **concentration** **stronger** acids



What makes for Acid Strength?

The strength of the acid is determined by the **number** of the **conjugate base** for the **acid**, when the **acid dissociates**: **The stronger the acid, the fewer** **conjugate base**



Water is **weaker** acid than **strong** acids, since **fewer** **conjugate base** (OH^-)

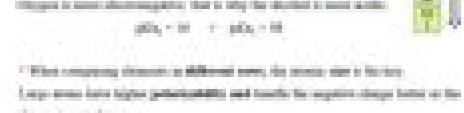
As **water** dissociates to **stronger** acids (H_3O^+), **stronger** acids (H_3O^+) are **stronger acids**

Water is **stronger** acid than **weak** acids (H_3O^+)

Acids with **higher** **concentration** are **stronger acids**

Acids with **lower** **concentration** are **weaker acids**

What does **stronger** dissociate at **different** **rate**, the **stronger** the **acid**.



Strong acids are **stronger** acids than **weak** acids

Weak acids are **stronger** acids than **strong** acids

What does **stronger** dissociate at **different** **rate**, the **stronger** the **acid**.



Strong acids are **stronger** acids than **weak** acids

Weak acids are **stronger** acids than **strong** acids

What does **stronger** dissociate at **different** **rate**, the **stronger** the **acid**.



Strong acids are **stronger** acids than **weak** acids

Weak acids are **stronger** acids than **strong** acids

What does **stronger** dissociate at **different** **rate**, the **stronger** the **acid**.

It has been shown the dissociating for acidity is called **Acid-base**.

Acid-base is **reversible** effect is the **transformation** of **strong** through **weak**.



It is more **electronegative** to **strong** **electrolyte** **weak** **electrolyte** the **stronger** the **acid**.



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Acid Base Reaction Study Guide

Susan McMurry

Acid Base Reaction Study Guide:

Organic Chemistry, 12e Binder Ready Version Study Guide & Student Solutions Manual T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2016-04-11 This is the Student Study Guide Solutions Manual to accompany Organic Chemistry 12th Edition The 12th edition of Organic Chemistry continues Solomons Fryhle Snyder's tradition of excellence in teaching and preparing students for success in the organic classroom and beyond A central theme of the authors approach to organic chemistry is to emphasize the relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us *Organic Chemistry, 13e Student Study Guide and Solutions Manual* T. W. Graham

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