

Acids and Bases

Overview of Acids and Bases

Acid-base Acids and Bases

Acids: Donates its proton to form conjugate base
 Bases: Accepts its proton to form conjugate acid



Remember: Only acids can donate, only bases can accept

What does conjugate acid/base and proton acceptor/donor?

Conjugate Acids and Bases

Acids are proton donors, bases are proton acceptors. The molecule is losing a proton is the acid, the molecule is gaining a proton is the base.



General Acids and Bases

Acid: Donates its proton to form conjugate base
 Base: Accepts its proton to form conjugate acid

Remember: Acid always donates protons, base always accepts protons

Acid Strength and pKa

Acid strength is measured by pKa
 Lower pKa = stronger acid



Remember: Lower pKa means stronger acid

As a general rule, stronger acids have lower pKa values

Acid Strength

What does a lower 'stronger acid'?

Remember: Lower pKa means stronger acid



The dissociation of the acid is a reversible reaction that is described by the equilibrium constant



Remember: $K_a \times K_b = K_w$

Acid Strength and pKa

The Ka value of weak acids (Ka) and weak bases (Kb) are shown below



The strength of a stronger acid/bases in the 4 figures

To compare the strength, $\log K_a$ is used, which is the pKa of the given acid

$$\text{pKa} = -\log K_a$$

$$\text{pKa}_1 + \text{pKa}_2 = 14 \quad \text{pKa}_1 + \text{pKa}_2 = 14$$

Remember: The stronger the acid, the lower the pKa value

Remember: The stronger the base, the higher the pKa value



What makes the Acid Strong?

The strength of the acid is determined by the stability of the conjugate base for the acid

The stronger the acid, the more stable its conjugate base



Remember: Stronger acids have stronger conjugate bases

Remember: The stronger the acid, the more stable its conjugate base

1) The first and best factor is the **inductive effect** which the backbone is connected, mainly to electronegative atoms (halogens and electronegative groups)

2) When the same atom, electronegative atoms that change such as electronegativity in the backbone



Remember: Stronger acids have stronger conjugate bases

$$\text{pKa}_1 + \text{pKa}_2 = 14$$

Remember: Stronger acids have stronger conjugate bases

Remember: Lower pKa means stronger acid



Remember: Stronger acids have stronger conjugate bases

The stronger the acid, the more stable its conjugate base

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The Position of Equilibrium for an Acid-Base Reaction

Remember: Stronger acids have stronger conjugate bases



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Choosing a Base to Deprotonate a Given Compound

Remember: Stronger acids have stronger conjugate bases

The stronger the acid, the more stable its conjugate base

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Choosing an Acid to Protonate a Given Compound

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The stronger the acid, the more stable its conjugate base

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Acid Base Reaction Study Guide

Susan McMurry



Acid Base Reaction Study Guide:

Organic Chemistry, 12e Binder Ready Version Study Guide & Student Solutions Manual T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2016-04-11 This is the Student Study Guide Solutions Manual to accompany Organic Chemistry 12th Edition The 12th edition of Organic Chemistry continues Solomons Fryhle Snyder s tradition of excellence in teaching and preparing students for success in the organic classroom and beyond A central theme of the authors approach to organic chemistry is to emphasize the relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us

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Study Guide [to Accompany] General Chemistry James E. Brady,1982

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