

25. The number of σ bonds in $\text{N}\equiv\text{N}$ is

- a. 1
- b. 2
- c. 3
- d. 4

26. The elements in an ionic compound are held together by

- a. electrostatic forces of attraction.
- b. van der Waals forces
- c. the spin of paired electrons.
- d. the formation of hybrid orbitals.
- e. an electron pair.

27. In every electrolytic and galvanic (voltaic) cell the anode is that electrode

- a. at which oxidation occurs.
- b. which attracts cations.
- c. at which electrons are supplied to the solution.
- d. at which reduction occurs.

28. Metal X was plated from a solution containing cations of X. The passage of 48.25 C deposited 31 mg of X on the cathode. What is the mass of X (in grams) per mole of electrons?

- a. 47
- b. 62
- c. 93
- d. 186

29. In a galvanic (voltaic) cell in which the reaction is $\text{Cd} + \text{Cu}^{2+} \rightarrow \text{Cu} + \text{Cd}^{2+}$ and the ions are at unit concentration (activity), the cell potential is.



30. In which reaction will an increase in total pressure at constant temperature favor formation of the products?

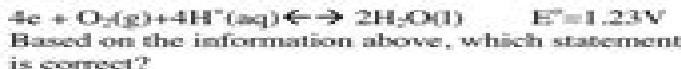
- a. $\text{CaCO}_3(s) \rightleftharpoons \text{CaO}(s) + \text{CO}_2(g)$
- b. $\text{H}_2(g) + \text{Cl}_2(g) \rightleftharpoons 2\text{HCl}(g)$
- c. $2\text{NO}(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}_2(g)$
- d. $\text{COCl}_2(g) \rightleftharpoons \text{CO}(g) + \text{Cl}_2(g)$

Standard Potentials	E°
$\text{Mg} \rightarrow \text{Mg}^{2+} + 2e^-$	2.37 V
$\text{Al} \rightarrow \text{Al}^{3+} + 3e^-$	1.66 V
$\text{Zn} \rightarrow \text{Zn}^{2+} + 2e^-$	0.76 V
$\text{Fe} \rightarrow \text{Fe}^{2+} + 2e^-$	0.44 V
$\text{Cu} \rightarrow \text{Cu}^{2+} + 2e^-$	0.34 V
$\text{Ag} \rightarrow \text{Ag}^{+} + e^-$	-0.80 V

31. Using only the metals Mg, Al, Zn, Fe, Cu and Ag, together with their 1 M salt solutions, a voltaic cell of the highest possible voltage would be constructed using electrodes of these metals.

- a. Mg and Ag
- b. Mg and Fe
- c. Zn and Cu
- d. Al and Ag
- e. Mg and Al

32. $E = E^\circ - 0.059/n \log Q$ (Nernst equation)
 $[\text{H}^+] = 1.0 \text{ M}$ initially, $P_{\text{O}_2} = 1.0 \text{ atm}$



- a. $n = 1$, since one mole of oxygen is being considered.
- b. Addition of base should result in an E value, which is less than 1.23 V.
- c. E is independent of the pH of the solution.
- d. $Q = \frac{[\text{H}_2\text{O}]}{[\text{O}_2][\text{H}^+]}$

33. The equilibrium constant for the gaseous reaction $\text{C} + \text{D} \rightleftharpoons \text{E} + \text{F}$ is 3.0 at 50 °C. In a 2.0 L flask at 50 °C are placed 1.0 mol of C, 1.0 mol of D, 1.0 mol of E, and 3.0 mol of F. Initially, the reaction will

- a. proceed at equal rates in both directions.
- b. proceed more rapidly to form E and F.
- c. proceed more rapidly to form C and D.
- d. not occur in either direction.

Compound	$\Delta G^\circ_f, \text{ kJ/mol}$
$\text{H}_2\text{O}(\text{l})$	-237
$\text{H}_2\text{O}(\text{g})$	-229

34. At 298 K the equilibrium constant for $\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightleftharpoons \text{H}_2\text{O}(\text{l})$

- a. is larger than the K_{eq} for $\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightleftharpoons \text{H}_2\text{O}(\text{g})$.
- b. will have a value of 1.0 at equilibrium.
- c. cannot be computed since data on O_2 and

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Norbert J. Pienta, Melanie M. Cooper, Thomas J. Greenbowe, 2005 Part of the Prentice Hall Series in Educational Innovation for Chemistry this unique book is a collection of information examples and references on learning theory teaching methods and pedagogical issues related to teaching chemistry to college students In the last several years there has been considerable activity and research in chemical education and the materials in this book integrate the latest developments in chemistry Each chapter is written by a chemist who has some expertise in the specific technique discussed has done some research on the technique and has applied the technique in a chemistry course *Abstracts of Papers American Chemical Society. Meeting*, *Griswold Chem Im 6/E* Nebergall, 1980 *Book of Abstracts American Chemical Society. Meeting*, American Chemical Society, 1987

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