

12: Stoichiometry

Key Stoichiometry Terms

- Stoichiometry:** Using the mole ratio in the balanced equation and information about one compound to find information about another in the reaction.
- Dimensional Analysis:** Method of converting units by multiplying by ratio of equalities.
- Molar Mass:** Sum of all the atomic masses (from the periodic table) in the compound.
- Solute:** Substance being dissolved in a homogeneous mixture (solution).
- Concentration:** Measure of how much solute is dissolved in how much solution.
- Molarity (M):** A concentration unit in moles per unit liter.
- Standard Temperature and Pressure (STP):** 1 atm (or 101.3 kPa) and 273 K (0°C).
- Molar volume of a gas:** 1 mole of any gas at STP is 22.4 liters.
- Limiting Reactant:** Reactant that stops the reaction by running out first.
- Actual Yield:** The amount actually produced in the lab.
- Theoretical Yield:** Amount that should be theoretically produced based on stoichiometric calculations.
- Percent Yield:** Compares the actual yield to the theoretical yield (their ratio with %).

Chemical Equations

- Coefficients give the mole ratios.

Example:



2 moles of H₂ react with 1 mole of O₂ and produces 2 moles of H₂O.

Dimensional Analysis in Stoichiometry

Equalities used during dimensional analysis for stoichiometry:

- Mole ratio in balanced equation:** Use to convert between moles of different compounds in the balanced equation.
- Molar mass:** Used to convert between grams and moles.
- Concentration:** Used to convert between moles and liters of a solution.
- Molar volume of a gas:** Used to convert between moles and liters of a gas at STP.

KUDOS for Stoichiometric Problems

Method to solve word problems:

- K = Known:** Identify all the known information (given quantities, balanced equation, etc.).
- U = Unknown:** What quantity and what compound does the problem ask for?
- D = Definition:** Identify the equalities that will be needed during dimensional analysis to convert from the known to the unknown.
- O = Output:** Perform the dimensional analysis.
- S = Substantiation:** Check for reasonableness, units and correct significant figures.

Molarity

Molarity - the concentration measure by the moles of solute per liter of solution.

$$\text{Molarity} = \frac{\text{moles solute}}{\text{L solution}}$$

Limiting Reactants

- Once a reactant has run out, the reaction will stop.
- Do stoichiometry for each given reactant quantity to the same product each time. Choose the calculation that gives the **smallest amount** of product.
- The reactant that produced the **smallest amount** of product is the **limiting reactant**.
- Steps to determine the limiting reactant: grams reactant \rightarrow moles reactant \rightarrow moles product \rightarrow grams product. Alternatively, use the molar method below.

Simple Limiting Reagent Finder Mnemonic: Find the moles of each reactant and divide the moles of each reactant by its coefficient. The reactant with the **smallest number** is the limiting reagent = "Smallest m/c is limited".

Percent Yield

$$\% \text{ yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

Example:

If stoichiometry calculations tell you that a reaction should produce 10.5 g, but when you perform the lab you produce 8.7 g, what was the percent yield?

$$\% \text{ yield} = \frac{8.7 \text{ g}}{10.5 \text{ g}} \times 100 = 82.9\%$$

Stoichiometry Examples

All examples for the following reaction:



Mole-Mass examples:

0.75 moles HCl react.	How many grams Mg are needed?
0.75 moles HCl	1 mole Mg $\frac{24.31 \text{ g Mg}}{2 \text{ mole HCl}}$ $= 12.15 \text{ g Mg}$

Mass-Mass example:

17.25 g Mg react.	How many grams MgCl ₂ are produced?
17.25 g Mg	1 mole Mg $\frac{95.21 \text{ g MgCl}_2}{1 \text{ mole Mg}}$ $= 95.21 \text{ g MgCl}_2$

Mass-Liter example (solution):

How many liters of 1.7 M HCl are needed to react with 2.5 g Mg?	2.5 g Mg	1 mole Mg	2 mole HCl	1 L HCl
	2.5 g Mg	1 mole Mg	$\frac{1.7 \text{ mole HCl}}{1 \text{ mole Mg}}$	$= 0.12 \text{ L HCl}$

Mass-Liter example (gas):

How many liters of H ₂ are produced when 3.5 g Mg react?	3.5 g Mg	1 mole Mg	1 mole H ₂	22.4 L H ₂
	3.5 g Mg	1 mole Mg	$\frac{1 \text{ mole H}_2}{1 \text{ mole Mg}}$	$= 3.5 \text{ L H}_2$

Limiting reactants:

If 0.25 mole HCl reacts with 0.55 mole Mg, how many grams of MgCl ₂ are produced and what was the limiting reactant?	0.25 mole HCl	1 mole Mg	$\frac{95.21 \text{ g MgCl}_2}{1 \text{ mole Mg}}$	95.21 g MgCl ₂
	0.25 mole HCl	2 mole Mg	$\frac{1 \text{ mole MgCl}_2}{2 \text{ mole Mg}}$	(none)

0.55 mole Mg	1 mole Mg	$\frac{95.21 \text{ g MgCl}_2}{1 \text{ mole Mg}}$	95.21 g MgCl ₂
	1 mole Mg	$\frac{0.55 \text{ mole MgCl}_2}{1 \text{ mole Mg}}$	52.6 g MgCl ₂

1.2 g MgCl₂ will be produced and HCl is the limiting reactant.

How to Use This Cheat Sheet: These are the keys related to this topic. Try to read through it carefully twice then write it out from memory on a blank sheet of paper. Review it again before the exam.

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