

1. According to the Brønsted-Lowry definition, which of the following species can function both as an acid and as a base?

- a. CH_4
- b. HCO_3^-
- c. CH_3CO_2^-
- d. H_2CO_3
- e. H_2O

2. In the reaction $\text{Ca(OH)}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{CaSO}_4 + \text{H}_2\text{O}$, which are the acids-base conjugate pair?

- a. H_2O and H_2O^-
- b. H_2O and H_2O^+
- c. Ca^{2+} and H_2O
- d. H_2SO_4 and SO_4^{2-}

3. Given that H_2O is a stronger Brønsted acid than H_2S in aqueous solution, which is true of a 1 M solution of NaHS ?

- a. It is more basic than a 1 M solution of NaH_2O .
- b. It is more basic than a 1 M solution of NaH_2S .
- c. It yields a neutral solution.
- d. It is more concentrated than a 1 M solution of NaH_2O .

4. The pH of a 0.01 M HCl solution is

- a. 1.5
- b. 2.0
- c. 3.0
- d. 5.0

5. The pH of a 1.0×10^{-3} M Ba(OH)_2 solution at 25°C is

- a. 2.0
- b. 2.5
- c. 11.0
- d. 11.5

6. Which is the strongest acid?

- a. HClO_4
- b. HClO_3
- c. HClO_2
- d. HClO_1

7. Which salt reacts with water (hydration) to generate a basic solution?

- a. $\text{Na}_2\text{CO}_3\text{H}_2\text{O}$
- b. Na_2SO_4
- c. NaClO_4
- d. $\text{Na}_2\text{SO}_4\text{H}_2\text{O}$

8. The position of the first ionization is

Acid	Conjugate Base	K_a (Strength of Acid)
HClO_4	ClO_4^-	0.0001% ionized
H_2SO_4	SO_4^{2-}	1.7×10^{-2}
H_3PO_4	PO_4^{3-}	5.0×10^{-3}
H_2O	H_3O^+	1.0×10^{-14}

- a. H_2O
- b. HClO_4
- c. H_3O^+
- d. SO_4^{2-}

Acid	Conjugate Base	K_a (Strength of Acid)
H_3PO_4	PO_4^{3-}	5.0×10^{-3}
H_2SO_4	SO_4^{2-}	1.7×10^{-2}
HClO_4	ClO_4^-	0.0001% ionized
H_2O	H_3O^+	1.0×10^{-14}

9. Which statement is a logical extension from the fact that a 0.1 M solution of potassium acetate KCH_3COO , is less acidic than a 0.1 M solution of potassium cyanide, KCN ?

- a. Acetate ion is a weaker acid than cyanide ion.
- b. Acetate ion is less soluble in water than cyanide ion.
- c. Cyanide ion is less soluble than acetate ion.
- d. Acetate ion is a stronger acid than hydrocyanic acid.

10. In the situation, will this salt, and its H_2O ionization, act as a strong acid, weak acid, or a strong base? Predict the proportionality, and the H_2O ionization, of the conjugate acid, with corresponding inversely to the proportionality of

- a. a 0.1 M sodium solution.
- b. a 0.01 M sodium hydroxide solution.
- c. a 0.001 M potassium nitrate solution.
- d. a 0.0001 M potassium bromate solution.

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