

# Acids & Bases Study Guide

6. Water is considered neutral. Describe what this means in terms of  $[H^+]$  &  $[OH^-]$  concentrations.

8. Complete the following:

$$pH = -\log[H^+]$$

$$K_w = [H^+][OH^-]$$

7. Fill in the blank:

9. What is the  $[OH^-]$  of a solution with a pH of 10?

10. What is the  $[H^+]$  of a solution with a pH of 4?

11. What is the pH of a solution with a  $[H^+]$  of  $1 \times 10^{-5}$  M?

12. What is the pH of a solution with a  $[OH^-]$  of  $1 \times 10^{-4}$  M?

13. What is the pH of a solution with a  $[H^+]$  of  $1 \times 10^{-12}$  M?

14. What is the pH of a solution with a  $[OH^-]$  of  $1 \times 10^{-10}$  M?

15. What is the pH of a solution with a  $[H^+]$  of  $1 \times 10^{-11}$  M?

16. What is the pH of a solution with a  $[OH^-]$  of  $1 \times 10^{-13}$  M?

Name: \_\_\_\_\_ Period: \_\_\_\_\_ Date: \_\_\_\_\_

## Acids & Bases Study Guide

1. Fill in the blanks or underline and contrast the properties of Acids & Bases.

Acids	Bases
<ul style="list-style-type: none"> <li>• Corrosive _____</li> <li>• Protons _____</li> <li>• Tastes _____</li> <li>• Turn litmus paper _____</li> <li>• High <math>[H^+]</math> _____</li> </ul>	<ul style="list-style-type: none"> <li>• Corrosive _____</li> <li>• Protons _____</li> <li>• Tastes _____</li> <li>• Turn litmus paper _____</li> <li>• High <math>[OH^-]</math> _____</li> </ul>

2. Fill in the chart for the following acids & bases.

Acid or Base?	Strong or Weak?	Formula	Name
			hydrochloric acid
			barium hydroxide
		$HClO_4$	
		$NaOH$	
		$H_2SO_4$	hydrochloric acid

3. Predict the products for the following double replacement reactions involving acids & bases. Identify which component is the Acid, Base, Conjugate Acid (CA), and Conjugate Base (CB) in the formula.



4. Label the pH Scale below (Strong Acid, Weak Acid, Base, Neutral, Strong Base, Weak Base, Salt, Soap).



# Acids And Bases Unit Study Guide

**John Krenos, Joseph Potenza, Loretta  
Jones, Lynn Koplitz, Thomas Spence**



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