

Name:

Block:

Solutions, Acids, and Bases Test Review

Solutions

- 1- In soda pop, $\text{CO}_2(\text{g})$ is a **solute** and water is the **solvent**.
- 2- What term is used by chemists to quantitatively describe a solution in which a relatively small amount of solute is dissolved? **Dilute/unsaturated**
- 3- When a solvent has dissolved all the solute it can at a particular temperature, it is said to be _____. **Saturated**
- 4- What will increase the rate of dissolution of a solute in a solvent? **Stirring, increasing surface area, increasing temperature**
- 5- What is the mass of H_2SO_4 in 1.00×10^2 mL of 0.200 M H_2SO_4 solution? **1.96g**
- 6- What number of moles of solute are present in 25.0 mL of 2.00 M HCl? **0.05 Moles**
- 7- How many grams of K_2SO_4 (molar mass = 174 g/mol) would be needed to prepare 4.00 L of a 0.0510 M solution? **35.5g**
- 8- A solution is prepared by dissolving 7.31 g of Na_2SO_4 in enough water to make 225 mL of solution. Calculate the solution molarity. **0.229M**
- 9- A 54.8 g sample of SrCl_2 is dissolved in 112.5 mL of solution. Calculate the molarity of this solution. **3.07M**

Gas Laws

- 1- A balloon has a volume of 1.20 L at 24.0°C . The balloon is heated to 48.0°C . Calculate the new volume of the balloon. **1.30L**
- 2- A sample of helium gas occupies 2.65 L at 1.20 atm. What pressure would this sample of gas exert in a 1.50 L container at the same temperature? **2.12atm**
- 3- A gas has a volume of 5.0 L at a certain pressure. How must the pressure be changed to double the volume of the gas at constant temperature? **Pressure is 1/2**
- 4- A 6.5 L sample of nitrogen at 25°C and 1.5 atm is allowed to expand to 13.0 L. The temperature remains constant. What is the final pressure? **0.75atm**
- 5- You transfer a sample of a gas at 17°C from a volume of 5.67 L and 1.10 atm to a container at 37°C that has a pressure of 1.10 atm. What is the new volume of the gas? **6.06L**
- 6- A sample of helium gas occupies 12.4 L at 23°C and 0.956 atm. What volume will it occupy at 40°C and 0.956 atm? **13.1L**
- 7- A sample of oxygen gas (O_2) has a volume of 7.51 L at a temperature of 19°C and a pressure of 1.38 atm. Calculate the moles of O_2 molecules present in this gas sample. **0.433 moles**
- 8- A weather balloon contains helium gas at a volume of 855 L. Initially, the balloon is at sea level, where the temperature is 25°C and the barometric pressure is 0.947 atm. The balloon then rises to an altitude of 6000 ft, where the pressure is 0.796 atm and the temperature is 15°C . What is the change in the volume of the balloon as it ascends from sea level to 6000 ft? **128L**
- 9- A 6.35 L sample of carbon monoxide is collected at 55°C and 0.892 atm. What volume will the gas occupy at 1.05 atm and $20.^\circ\text{C}$? **4.82L**

Acids And Bases Test Review Answers

Ruth E. McCall



Acids And Bases Test Review Anser:

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Study Guide to accompany Basic Concepts of Chemistry, 7th Edition Leo J. Malone, 2003-02-20. Work more effectively and gauge your progress along the way. This Study Guide that is designed to accompany Malone's Chemistry 7th Edition includes chapter summaries, new terms, self tests, answers to self tests, and solutions to selected problems. This easy to read introduction presents chemistry as a living relevant science. Chemistry 7th Edition encourages critical thinking and helps readers overcome the math difficulties that often prevent them from developing a full understanding of the subject. This new seventh edition builds on its core strengths of pedagogy, driving the connections between ideas, mathematics in context, not just an appendix, and an extensive problem solving emphasis with an updated design and more molecular art. In addition, the seventh edition expands its applications and online options. One of the briefest books in the market, it still provides sufficient depth for the basic concepts of chemistry. **The Electrical Review**, 1876 [Iron Trade Review and Western](#)

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