

Exam #10 Review: Acids, Bases, and pH

1. Know the properties of acids and bases. What are their similarities and differences?

Acids

- Taste sour
- React with metals to form hydrogen gas
- Contain a hydrogen ion (H^{+1})
- pH value less than 7

Bases

- Taste bitter
- Feel slippery
- Contain a hydroxide ion (OH^{-1})
- pH value greater than 7

BOTH

- change colors of indicators
- react with each other to form salt and water
- conduct electricity when dissolved in solution (electrolytes)

2. Know the rules for naming **acids** (binary acid and oxyacid) and name the following below:

Binary acids (only 2 elements): hydro – element – ic acid

Oxyacids (contain a polyatomic ion): NO "HYDRO" IN THE NAME

- -ate ion \rightarrow -ic acid
- -ite ion \rightarrow -ous acid

a. HCl hydrochloric acid
b. HF hydrofluoric acid
c. $HClO_2$ chlorous acid

d. HNO_3 nitric acid
e. H_3PO_4 phosphoric acid
f. H_2SO_3 sulfurous acid

REMEMBER:
"I ate something icky,
all nite I was nauseous"

3. Know the rules for naming **bases** and name the following below:

Element + hydroxide

a. NaOH sodium hydroxide
b. LiOH lithium hydroxide

c. $Al(OH)_3$ aluminum hydroxide
d. $Ba(OH)_2$ barium hydroxide

4. How are Arrhenius acids and bases defined?

Acids – produce H^{+1} ions in aqueous solutions

Bases – produce OH^{-1} ions in aqueous solutions

5. How are Bronsted-Lowry acids and bases defined?

Acids – donate (give away) a H^{+1} ion (proton)

Bases – accept (take in) a H^{+1} ion (proton)

6. Show how the following acids and bases ionize:



7. If it is a strong acid or base, how much do they ionize in solution?

STRONG acids & bases **COMPLETELY** ionize (all break apart) in a solution

8. If it is a weak acid or base, how much do they ionize in solution?

WEAK acids & bases do **NOT** completely ionize

9. What is a neutralization reaction? What products are always produced?

Acid reacts with a base always produces water and a salt

10. Complete the following neutralization reaction and what type of salt is made:



Type of Salt: neutral



Type of Salt: neutral

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James H. Halsey

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