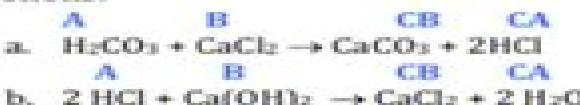


Chapter 19 Study Guide: Acids, Bases & Titration – Answers

1. Differentiate between Arrhenius, Brønsted-Lowry & Lewis acids and bases.

	ACIDS	BASES
Arrhenius	Form hydronium ions (H_3O^+) in aqueous sol'n	Form hydroxide ions (OH^-) in aqueous sol'n
Brønsted-Lowry	Proton (H^+) donors	Proton (H^+) acceptors
Lewis	Electron pair acceptors	Electron pair donors

2. Label the acid (A), base (B), conjugate acid (CA) and conjugate base (CB) in each of the following reactions:



3. Write the names or formulas for the following acids and bases:

a. H_2S – hydrosulfuric acid	e. Nitrous acid – HNO_2
b. $\text{HC}_2\text{H}_3\text{O}_2$ – acetic acid	f. Calcium hydroxide – $\text{Ca}(\text{OH})_2$
c. NaOH – sodium hydroxide	g. hydrochloric acid – HCl
d. HNO_3 – nitric acid	

Use the following formulas for the remaining questions: (these will not be given to you on the test!)

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$$

$$\text{pH} = -\log[\text{H}_3\text{O}^+]$$

$$\text{pOH} = -\log[\text{OH}^-]$$

$$\text{pH} + \text{pOH} = 14$$

$$\text{M}_\text{A}\text{-V}_\text{A} \cdot \text{n}_\text{A} = \text{M}_\text{B}\text{-V}_\text{B} \cdot \text{n}_\text{B}$$

4. What is the pH of 1.34×10^{-4} M solution of hydrochloric acid?

$$\text{pH} = -\log[1.34 \times 10^{-4}\text{M}] = 3.873$$

5. Determine the pH and pOH of a 1.85×10^{-6} M NaOH solution.

$$\text{pOH} = -\log[1.85 \times 10^{-6}\text{M}] = 5.733$$

$$\text{pH} = 14 - 5.73 = 8.267$$

6. What is the $[\text{OH}^-]$ of an aqueous solution that has a pH of 5.0?

$$\text{pOH} = 9 \quad [\text{OH}^-] = 1 \times 10^{-9}\text{M}$$

7. What is the pH and the pOH of a solution with a volume of 2000. mL containing 35 grams of NaOH? Is this solution acidic or basic?

$$\frac{35\text{gNaOH}}{2.000\text{L}} \times \frac{1\text{molNaOH}}{40.00\text{gNaOH}} = 0.44\text{M NaOH}$$

$$\text{pOH} = -\log[0.44\text{M}] = 0.36 \quad \text{pH} = 14 - 0.36 = 13.64 \quad \text{basic}$$

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David Margolese,David W. Oxtoby

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The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills,Amina El-Ashmawy,2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone,Theodore O.

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