

Answers

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TASK 1 — Bronsted-Lowry acids & bases

| | | |
|--|---|---|
| 1. Acid = H_2O , base = NH_3 | 2. Acid = HCl , base = H_2O | 3. Acid = HCOOH , base = KOH |
| 4. Acid = HCl , base = CH_3COOH | 5. Acid = HCl , base = NH_3 | 6. Acid = HCO_3^- , base = OH^- |
| 7. Acid = H^+ , base = HCO_3^- | 8. Acid = H_2SO_4 , base = HNO_3 | |

TASK 2 — pH of strong acids

| | | | |
|-----------------------------|-----------|-----------|-------------------------------|
| 1. a) 0.70 | b) 1.30 | c) 1.10 | d) -0.30 |
| 2. a) 2.82×10^{-2} | b) 0.0100 | c) 0.0501 | d) 1.58 mol dm^{-3} |
| 3. a) 1.46 | b) 0.80 | c) 0.10 | d) 0.62 |
| 4. a) 0.56 | b) 0.39 | c) 0.10 | d) -0.31 |

| | | |
|----|---|--|
| 1. | a) $\text{pH} = -\log 0.2 = 0.70$ | |
| | b) $\text{pH} = -\log 0.05 = 1.30$ | |
| | c) $\text{pH} = -\log 0.001 = 1.10$ | |
| | d) $\text{pH} = -\log 0.0001 = -0.30$ | |
| 2. | a) $\text{pH} = -\log 10^{-2.82} = 2.82 \times 10^{-2}$ | c) $\text{pH} = -\log 0.0501 = 1.30$ |
| | b) $\text{pH} = -\log 0.0100 = 2.00$ | d) $\text{pH} = -\log 0.0001 = 4.00$ |
| | c) $\text{pH} = -\log 0.001 = 3.00$ | e) $\text{pH} = -\log 0.00001 = 4.60$ |
| | d) $\text{pH} = -\log 0.1 = 0.10$ | f) $\text{pH} = -\log 0.000001 = 5.30$ |
| 3. | a) $\text{pH} = -\log 0.0001 = 4.00$ | $\text{pH} = -\log 0.000001 = 5.30$ |
| | b) $\text{pH} = -\log 0.00001 = 4.60$ | $\text{pH} = -\log 0.0000001 = 6.00$ |
| | c) $\text{pH} = -\log 0.001 = 3.00$ | $\text{pH} = -\log 0.0000001 = 6.60$ |
| | d) $\text{pH} = -\log 0.0001 = 3.40$ | $\text{pH} = -\log 0.0000001 = 7.00$ |
| 4. | a) $\text{pH} = -\log 0.2 = 0.70$ | $\text{pH} = -\log 0.0001 = 4.00$ |
| | b) $\text{pH} = -\log 0.0001 = 4.00$ | $\text{pH} = -\log 0.00001 = 4.60$ |
| | c) $\text{pH} = -\log 0.001 = 3.00$ | $\text{pH} = -\log 0.000001 = 5.30$ |
| | d) $\text{pH} = -\log 0.1 = 1.00$ | $\text{pH} = -\log 0.0000001 = 6.90$ |

Acid And Bases Chemistry Workbook Answers

David R. Klein, Laurie S. Starkey

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