

Section 3.1 Solids, Liquids, and Gases

Gases

- A *gas* is the state of matter that does not have a definite shape or a definite volume.

- Give three examples of gases:
 - 1.) Oxygen
 - 2.) Carbon Dioxide
 - 3.) Steam

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web model i types of reactions set a $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$ b $\text{Fe}_2\text{O}_3 + 3\text{C} \rightarrow 2\text{Fe} + 3\text{CO}$ c $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$ d $\text{Fe}_2\text{O}_3 + 3\text{H}_2 \rightarrow 2\text{Fe} + 3\text{H}_2\text{O}$ e $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$ f $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$ g $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$ h $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$ i $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$ j $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$ k $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$ l $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$ m $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$ n $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$ o $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$ p $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$ q $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$ r $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$ s $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$ t $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$ u $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$ v $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$ w $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$ x $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$ y $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$ z $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$

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