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web exam tip you need to appreciate that the measurement of amounts in moles can apply to atoms molecules ions electrons formulae and equations e g in one mole of carbon c the number of atoms is the same as the number of molecules in one mole of carbon dioxide co₂ linking the mole and the atomic mass

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web for practical purposes a rounded version of the constant can be used in exams 6.02×10^{23} mol⁻¹ the reason we have mol⁻¹ as the unit is because this is the number of entities per mole of substance for example one mole of sodium Na contains 6.02×10^{23} atoms of sodium one mole of hydrogen H₂ contains 6.02×10^{23} molecules of hydrogen