

Acids, Bases and Salts

- **THE THEORY of ACIDS and ALKALIS** and a few technical terms:
 - **Acids** are substances that **form hydrogen ions ($\text{H}^+_{(\text{aq})}$)** when dissolved in water eg hydrochloric acid HCl gives $\text{H}^+_{(\text{aq})}$ and $\text{Cl}^-_{(\text{aq})}$ ions, sulphuric acid H_2SO_4 gives $2\text{H}^+_{(\text{aq})}$ and SO_4^{2-} ions and nitric acid HNO_3 gives $\text{H}^+_{(\text{aq})}$ and $\text{NO}_3^-_{(\text{aq})}$ ions.
 - **Alkalis** are substances that **form hydroxide ions ($\text{OH}^-_{(\text{aq})}$)** in water eg sodium hydroxide NaOH gives $\text{Na}^+_{(\text{aq})}$ and $\text{OH}^-_{(\text{aq})}$ ions, calcium hydroxide $\text{Ca}(\text{OH})_2$ gives $\text{Ca}^{2+}_{(\text{aq})}$ and $2\text{OH}^-_{(\text{aq})}$ ions. Note: **an alkali is a base soluble in water.**
 - In **water**, there are trace quantities of H^+ and OH^- ions BUT they are of equal concentration and so water is neutral.
 - In **acid solutions** there are **more H^+ ions than OH^- ions.**
 - In **alkaline solution** there are **more OH^- ions than H^+ ions.**
 - Acids dissociate to different extents in aqueous solution. Acids that dissociate to a large extent are strong electrolytes and **strong acids**. In contrast, acids that dissociate only to a small extent are **weak acids** and weak electrolytes.

In a similar manner, bases can be strong or weak depending on the extent to which they dissociate and produce OH^- ions in solution. Most metal hydroxides are strong electrolytes and **strong bases**. Ammonia, NH_3 , is a weak electrolyte and **weak base**.



- **BASES** eg oxides and hydroxides are substances that **react and neutralise acids to form salts and water**. **Bases which are soluble in water are called alkalis.**

Acids

Some common acids are listed below:

Name	Formula	Strong/Weak	Where is it found?
Hydrochloric acid	HCl	Strong	The stomach, in the lab.
Sulphuric acid	H_2SO_4	Strong	Acid rain, car batteries, the lab.
Nitric acid	HNO_3	Strong	Acid rain, in the lab.
Ethanoic (acetic) acid	CH_3COOH	Weak	Vinegar
Methanoic (formic) acid	HCOOH	Weak	Ant & nettle stings, descalers
Citric Acid	$\text{C}_6\text{H}_8\text{O}_7$	Weak	Citrus fruits

- Acids taste sour (e.g. vinegar, lemon juice).
- Acids are harmful to living cells.
- Aqueous solutions of all acids contain **hydrogen ions, H^+** .

Acids Bases And Salts Study Guide

Alexander Silverman



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