

Name:

Key

Period:

Acids and Bases

How do acids and bases behave in water?

Why?

Acids and bases play an important role in our lives. Numerous biological processes, industrial applications, and even environmental problems are a function of the acidity or basicity (alkalinity) of aqueous solutions. It is therefore important to understand what makes a substance behave as an acid or a base when dissolved in water. In this activity, we will explore the physical and chemical properties of acids and bases.

Model 2 - Bronsted-Lowry Acids and Bases



1. Identify the Bronsted-Lowry acids in Model 2.

a. Atoms of which element are present in all of the Bronsted-Lowry acids in Model 2?

Hydrogen

b. How can you tell from Reaction 1 that HCl loses an H^+ ion rather than a hydrogen atom when the reaction occurs. Hint: Look at the products.

The Cl has a negative charge. The positive from the hydrogen and the negative from the chloride cancel each other out when in a compound.

2. For each acid-base reaction in Model 2, describe the role of the Bronsted-Lowry acid in the H^+ ion (proton) transfer that occurs.

The acids lose a hydrogen ion

Acids And Bases Pogil Extension Questions

Brian G. Cox

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