

Section 1 Acids

Goals

- To describe acids
- To make the distinction between strong and weak acids.
- To show the changes that take place on the particle level when acids dissolve in water.
- To show how you can recognize strong and weak acids.

The most important skills to develop in this section are to be able to recognize acids from names or formulas.

Identification of Acids and Bases: Identify each of the following as an Arrhenius strong acid, an Arrhenius weak acid, an Arrhenius strong base, or an Arrhenius weak base.

- HNO_3 **strong acid**
- lithium hydroxide **strong base**
- K_2CO_3 **weak base**
- hydrofluoric acid **weak acid**

Section 2 Acid Nomenclature

- *Goal: To describe how to convert between names and chemical formulas for acids.*

This section adds acids to the list of compounds for which you should be able to convert between names and formulas.

Formulas for Acids: Write the chemical formulas that correspond to the names

- hydrofluoric acid: This name has the form of a binary acid, hydro(root)ic acid, so its formula is HF(aq) .
- phosphoric acid: This name has the form of an oxyacid, (root)ic acid, so it contains hydrogen, phosphorus, and oxygen. Phosphate is PO_4^{3-} , so phosphoric acid is H_3PO_4 .

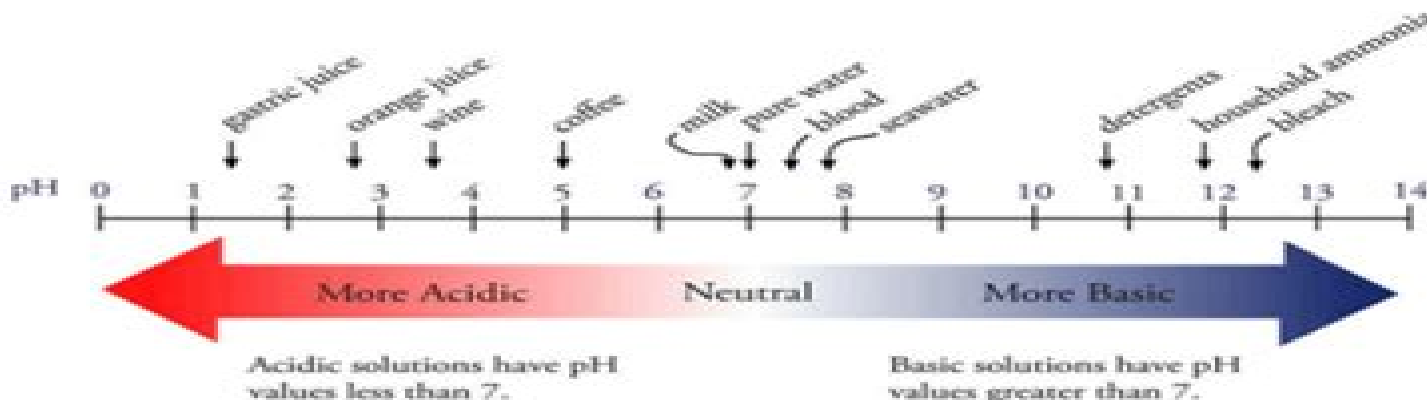
Naming Acids: Write the names that correspond to the chemical formulas

- HI : This is a binary acid, so its name has the form hydro(root)ic acid. HI(aq) is **hydroiodic acid**.
- $\text{HC}_2\text{H}_3\text{O}_2$: The name of the oxyanion $\text{C}_2\text{H}_3\text{O}_2^-$ is acetate, so $\text{HC}_2\text{H}_3\text{O}_2$ is **acetic acid**.

Section 3 Summary of Chemical Nomenclature

- *Goal: To review the process for converting between names and formulas for binary and ternary covalent, ionic compounds, acids, and oxyacids compounds.*

The ability to convert among names and formulas for chemical compounds is crucial to communication between chemists and chemistry students.



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