

## AP Biology 013 – Photosynthesis and Respiration

### Video Review Sheet

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1. What is the goal of life?
  - The goal of life is to make ATP so we can use energy.
2. Name and describe and give an example of the two life strategies.
  - The two life strategies are the autotrophs and heterotrophs.  
Autotrophs- make their own food. Ex. Plant  
Heterotroph- can't make their own. Ex. Humans
3. What are the two types of autotrophs?
  - Two types of autotrophs is Chemosynthesis and Photosynthesis
4. What are the two types of heterotrophs?
  - Two types of heterotrophs is Cellular Respiration and Fermentation
5. Who are the autotrophs making the food for?
  - Autotrophs making food from themselves.
6. Where does chemosynthesis occur?
  - Chemosynthesis occurs when there's not a lot of light on our planet.
7. Tubeworms contain what type of chemosynthetic organism?
  - Black smokers
8. Yeast can do alcoholic fermentation. Lactic acid fermentation, that we do, also does not require: **oxygen**.
9. Write the equation for photosynthesis (reverse is cellular respiration):
  - $6 \text{ CO}_2 + \text{H}_2\text{O} \rightarrow \text{C}_6\text{H}_{12}\text{O}_6 + 6 \text{ O}_2$
10. What are you storing in the glucose molecule?
  - Taking the carbon in carbon dioxide and we're actually turning that into carbon in glucose releasing oxygen as a waste product.
11. What is the delta G?
  - Positive free energies
12. Where does the light reaction take place specifically?
  - Light reaction takes place in the thylakoid membrane

# Bio Guide Answers Photosynthesis

**Martha R. Taylor**



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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web goals understand that parts of an acid base titration be able to determine the  $K_a$  or  $K_b$  from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate  $K_{a1}$  and  $K_{a2}$  for a polyprotic acid by the end of this lab students should be able to

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance  $\text{H}_2\text{C}_2\text{O}_4$  aq  $2\text{NaOH}$  aq  $\text{Na}_2\text{C}_2\text{O}_4$  4h20 to differentiate an acid from a base in a solution a ph scale is used the ph scale ranges from 0 to 14 a ph of 7 is neutral in an acid base reaction

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web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction  $\text{NaOH}$  aq  $\text{C}_6\text{H}_4\text{COOH}$  aq  $\text{C}_6\text{H}_4\text{COONa}$  aq  $\text{H}_2\text{O}$  l the point of equivalence is the point of titration when the number of moles of naoh na added is equal to the number of moles of khp nb in the

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