

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentrations from  $\text{pH}$  and  $\text{pOH}$ , and vice versa from  $\text{pH}$  and  $\text{pOH}$ .
3. Calculate the  $\text{pH}$  of a weak acid.

## Vocabulary

- |                   |                               |                  |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. $\text{pH}$                | 6. $\text{pOH}$  |

## Key Equation

- |                                     |                                  |
|-------------------------------------|----------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = 14 - \text{pOH}$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = 14 - \text{pOH}$ |

## Part A Completion

Write the hydrogen ion concentration or autoionization constant, if the missing quantity that is indicated in the partial equation below. Complete each equation as needed.

1. When autoionization is  $1.0 \times 10^{-14}$ , autoionization constant is  $1.0 \times 10^{-14}$ .
2. When autoionization is  $1.0 \times 10^{-14}$ , the autoionization constant is  $1.0 \times 10^{-14}$ .
3. When autoionization is  $1.0 \times 10^{-14}$ , the autoionization constant is  $1.0 \times 10^{-14}$ .
4. When autoionization is  $1.0 \times 10^{-14}$ , the autoionization constant is  $1.0 \times 10^{-14}$ .
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1.  $1.0 \times 10^{-14}$
2.  $1.0 \times 10^{-14}$
3.  $1.0 \times 10^{-14}$
4.  $1.0 \times 10^{-14}$
5.  $1.0 \times 10^{-14}$
6.  $1.0 \times 10^{-14}$
7.  $1.0 \times 10^{-14}$
8.  $1.0 \times 10^{-14}$
9.  $1.0 \times 10^{-14}$
10.  $1.0 \times 10^{-14}$

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