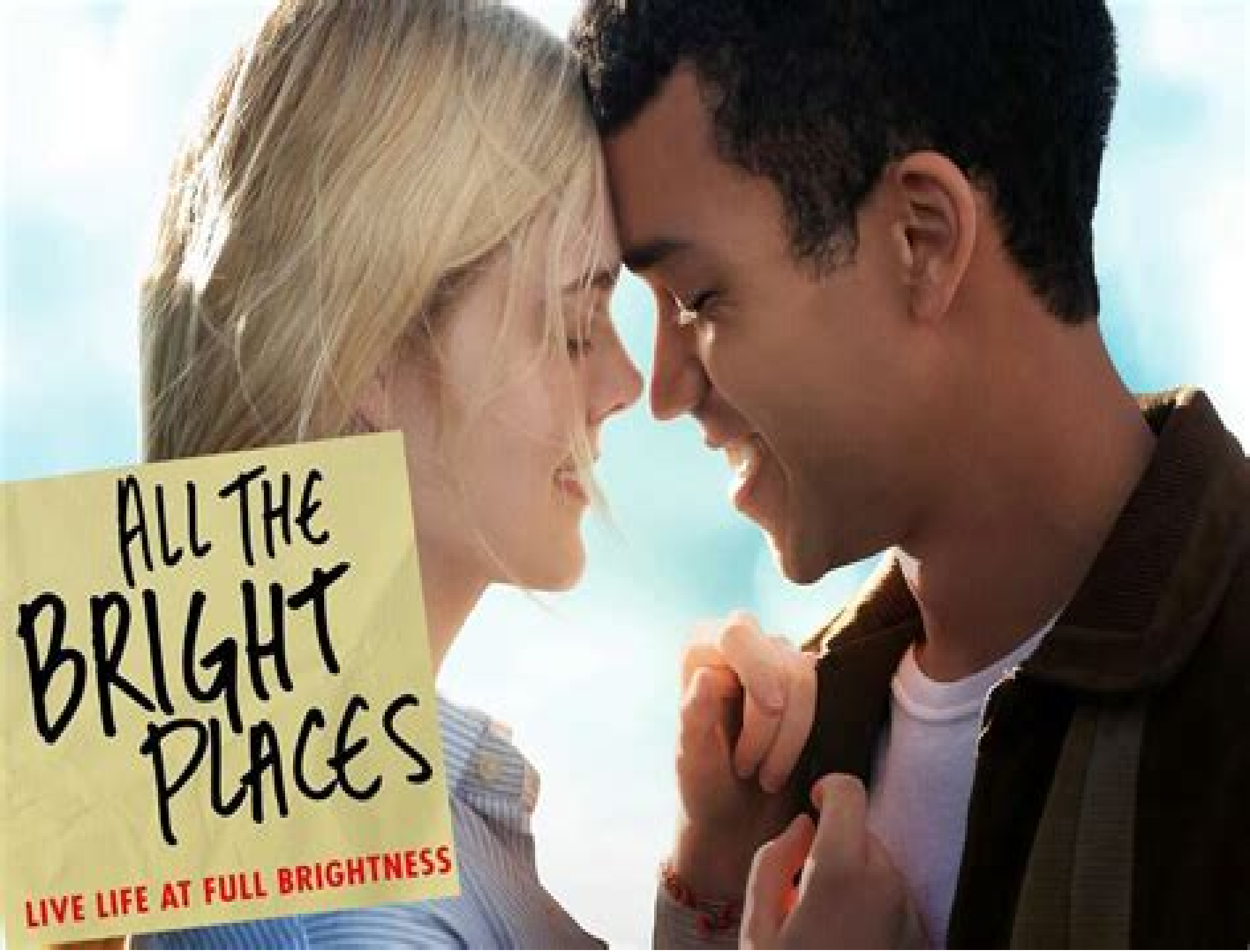




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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- + 2\text{Ag}^+ \rightarrow 2\text{AgCl}$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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$1\text{K}^+ + 3\text{PO}_4^{3-} + \text{Al}^{3+} + 3\text{OH}^- \rightarrow \text{AlPO}_4 + 3\text{OH}^-$ ionic $3\text{Cl}^- + \text{Al}^{3+} + 3\text{NH}_4^+ + 3\text{PO}_4^{3-} \rightarrow \text{AlPO}_4 + 3\text{NH}_4^+$ net ionic $\text{Al}^{3+} + \text{PO}_4^{3-} \rightarrow \text{AlPO}_4$ s 2 2 fec13 aq 3 mg s

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 13 2ag aq 2f aq 2nh₄ aq cr₂o₇²⁻ aq ag₂cr₂o₇ s 2nh₄ aq 2f aq because two nh₄ aq

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