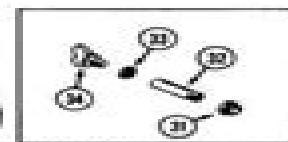
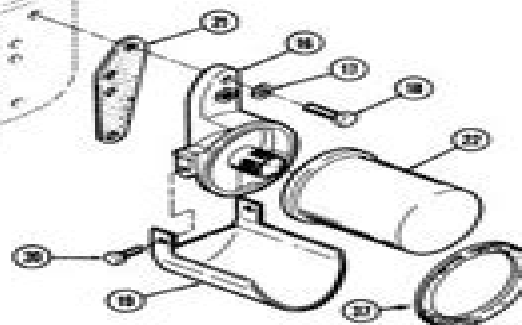


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Lingjun Ying



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web elements in the second row of the periodic table place their electrons in the 2n shell as well as the 1n shell for instance lithium li has three electrons two fill the 1 s orbital and the third is placed in the 2 s orbital giving an electron configuration of 1 s ² 2 s ¹

web electron configurations and trends answer these discussion questions given the following part of an electron configuration $3s^2$ explain which part refers to the number of electrons the energy level and the sublevel what is the maximum number of electrons that can occupy a single orbital state and explain hund s rule

web one more point needs to be emphasized about the relationship between electron configuration and the periodic table the atoms of elements in the same vertical column of the table have similar electron configurations for example consider the alkaline earth elements group iia

web the electron removed during the ionization of beryllium $1s^2 2s^2$ is an s electron whereas the electron removed during the ionization of boron $1s^2 2s^2 2p^1$ is a p electron this results in a lower first ionization energy for boron even though its nuclear charge is greater by one proton

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