

## Chemistry 2<sup>nd</sup> Semester Final Review KEY

\*This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.\*

Bring your book to school by Monday for turn in (seniors should bring book to bookstore by themselves ASAP)

Gas Laws (Ch. 13):

Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)

Practice questions:

1. If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)

$$V_1 = 3.8 \text{ L}; P_1 = 2.71 \text{ atm}$$

$$V_2 = 1.47 \text{ L}; P_2 = ? \text{ atm}$$

Since all units are standard (L, atm), we do not need to convert the units before proceeding.

$$(P_1 V_1) / T_1 = (P_2 V_2) / T_2$$

T goes away (not mentioned in problem)

$$P_1 V_1 = P_2 V_2$$

Isolate for P<sub>2</sub>:  $P_2 = (P_1 V_1) / V_2$

$$P_2 = (2.71 \text{ atm} \times 3.8 \text{ L}) / (1.47 \text{ L}) = 7.0 \text{ atm} \text{ (L cancelled and it's 2 sig figs)}$$

2. A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)

$$V_1 = 1.49 \text{ L}, T_1 = 34.75 \text{ degrees Celsius}$$

$$V_2 = ? \text{ L}, T_2 = 78.41 \text{ degrees Celsius}$$

We need to convert from Celsius to Kelvins before proceeding.

$$T_1 = 34.75 + 273 = 307.75 \text{ K}$$

$$T_2 = 78.41 + 273 = 351.41 \text{ K}$$

$$(P_1 V_1) / T_1 = (P_2 V_2) / T_2$$

Get rid of P

$$V_1 / T_1 = V_2 / T_2$$

Isolate V<sub>2</sub>:  $V_2 = (V_1 / T_1) T_2$

$$V_2 = (1.49 \text{ L} / 307.75 \text{ K}) (351.41 \text{ K}) = 1.70 \text{ L}$$

3. What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?

The combined gas law won't work here. You only have one condition and nothing changes.

PV=nRT will work. (n= moles and R= 0.08026 L atm/K mol)

P= 1.04 atm, V= ? L, n= OOPS we need to convert from grams to moles!, T= 84.7 degrees Celsius + 273= 357.7 K

n: Hydrogen gas is diatomic!

$$8.47 \text{ g H}_2 \times (1 \text{ mol H}_2 / 2.02 \text{ g H}_2) = 4.193069307 \text{ mol H}_2$$

$$\text{Isolate V: } V = (nRT) / P = (4.193069307 \text{ mol H}_2 \times 0.08206 \text{ L atm/ K mol} \times 357.7 \text{ K}) / 1.04 \text{ atm} = 118 \text{ L}$$

4. What volume is occupied by 56.75 g of oxygen gas at STP?

# 2nd Semester Chemistry Final Exam Review Answers

**Mary Millhollon, Richard H. Langley**

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## **Table of Contents 2nd Semester Chemistry Final Exam Review Answers**

1. Understanding the eBook 2nd Semester Chemistry Final Exam Review Answers
  - The Rise of Digital Reading 2nd Semester Chemistry Final Exam Review Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying 2nd Semester Chemistry Final Exam Review Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 2nd Semester Chemistry Final Exam Review Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from 2nd Semester Chemistry Final Exam Review Answers
  - Personalized Recommendations
  - 2nd Semester Chemistry Final Exam Review Answers User Reviews and Ratings
  - 2nd Semester Chemistry Final Exam Review Answers and Bestseller Lists
5. Accessing 2nd Semester Chemistry Final Exam Review Answers Free and Paid eBooks
  - 2nd Semester Chemistry Final Exam Review Answers Public Domain eBooks
  - 2nd Semester Chemistry Final Exam Review Answers eBook Subscription Services
  - 2nd Semester Chemistry Final Exam Review Answers Budget-Friendly Options

6. Navigating 2nd Semester Chemistry Final Exam Review Answers eBook Formats
  - ePUB, PDF, MOBI, and More
  - 2nd Semester Chemistry Final Exam Review Answers Compatibility with Devices
  - 2nd Semester Chemistry Final Exam Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 2nd Semester Chemistry Final Exam Review Answers
  - Highlighting and Note-Taking 2nd Semester Chemistry Final Exam Review Answers
  - Interactive Elements 2nd Semester Chemistry Final Exam Review Answers
8. Staying Engaged with 2nd Semester Chemistry Final Exam Review Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 2nd Semester Chemistry Final Exam Review Answers
9. Balancing eBooks and Physical Books 2nd Semester Chemistry Final Exam Review Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 2nd Semester Chemistry Final Exam Review Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine 2nd Semester Chemistry Final Exam Review Answers
  - Setting Reading Goals 2nd Semester Chemistry Final Exam Review Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 2nd Semester Chemistry Final Exam Review Answers
  - Fact-Checking eBook Content of 2nd Semester Chemistry Final Exam Review Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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