

25. The number of σ bonds in $N \equiv N$ is
- 1
 - 2
 - 3
 - 4
26. The elements in an ionic compound are held together by
- electrostatic forces of attraction.
 - van der Waals forces
 - the spin of paired electrons.
 - the formation of hybrid orbitals.
 - an electron pair.
27. In every electrolytic and galvanic (voltaic) cell the anode is that electrode
- at which oxidation occurs.
 - which attracts cations.
 - at which electrons are supplied to the solution.
 - at which reduction occurs.
28. Metal X was plated from a solution containing cations of X. The passage of 48.25 C deposited 31 mg of X on the cathode. What is the mass of X (in grams) per mole of electrons?
- 47
 - 62
 - 93
 - 186
29. In a galvanic (voltaic) cell in which the reaction is $Cd + Cu^{2+} \rightarrow Cu + Cd^{2+}$ and the ions are at unit concentration (activity), the cell potential is
- $$\begin{array}{ll} Cd \rightarrow Cd^{2+} + 2e^- & 0.4021 \text{ V} \\ Cu \rightarrow Cu^{2+} + 2e^- & -0.344 \text{ V} \end{array}$$
- 0.1383 V
 - 0.4021 V
 - 0.344 V
 - 0.7461 V
 - 0.3677 V
30. In which reaction will an increase in total pressure at constant temperature favor formation of the products?
- $CaCO_3(s) \rightleftharpoons CaO(s) + CO_2(g)$
 - $H_2(g) + Cl_2(g) \rightleftharpoons 2HCl(g)$
 - $2NO(g) + O_2(g) \rightleftharpoons 2NO_2(g)$
 - $COCl_2(g) \rightleftharpoons CO(g) + Cl_2(g)$

Standard Potentials	E°
$Mg \rightarrow Mg^{2+} + 2e^-$	2.37V
$Al \rightarrow Al^{3+} + 3e^-$	1.66V
$Zn \rightarrow Zn^{2+} + 2e^-$	0.76V
$Fe \rightarrow Fe^{2+} + 2e^-$	0.44V
$Cu \rightarrow Cu^{2+} + 2e^-$	0.34V
$Ag \rightarrow Ag^+ + e^-$	-0.80V

31. Using only the metals Mg, Al, Zn, Fe, Cu and Ag, together with their 1 M salt solutions, a voltaic cell of the highest possible voltage would be constructed using electrodes of these metals.
- Mg and Ag
 - Mg and Fe
 - Zn and Cu
 - Al and Ag
 - Mg and Al
32. $E = E^\circ - 0.059/n \log Q$ (Nernst equation)
 $[H^+] = 1.0 \text{ M}$ initially, $P_{O_2} = 1.0 \text{ atm}$
- $$4e^- + O_2(g) + 4H^+(aq) \rightleftharpoons 2H_2O(l) \quad E^\circ = 1.23 \text{ V}$$
- Based on the information above, which statement is correct?
- $n = 1$, since one mole of oxygen is being considered.
 - Addition of base should result in an E value, which is less than 1.23 V.
 - E is independent of the pH of the solution.
 - $Q = \frac{[H_2O]^2}{[O_2][H^+]}$
33. The equilibrium constant for the gaseous reaction $C + D \rightleftharpoons E + 2F$ is 3.0 at 50 °C. In a 2.0 L flask at 50 °C are placed 1.0 mol of C, 1.0 mol of D, 1.0 mol of E, and 3.0 mol of F. Initially, the reaction will
- proceed at equal rates in both directions.
 - proceed more rapidly to form E and F.
 - proceed more rapidly to form C and D.
 - not occur in either direction.
- | Compound | ΔG_f° , kJ/mol |
|-----------|-----------------------------|
| $H_2O(l)$ | -237 |
| $H_2O(g)$ | -229 |
34. At 298 K the equilibrium constant for $H_2(g) + \frac{1}{2} O_2(g) \rightleftharpoons H_2O(l)$
- is larger than the K_{eq} for $H_2(g) + \frac{1}{2} O_2(g) \rightleftharpoons H_2O(g)$
 - will have a value of 1.0 at equilibrium.
 - cannot be computed since data on O_2 and

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Diane M. Bunce, Cinzia M. Muzzi



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Broadening Participation in STEM Zayika Wilson-Kennedy,Goldie S. Byrd,Eugene Kennedy,Henry T. Frierson,2019-02-28 This book reports on high impact educational practices and programs that have been demonstrated to be effective at broadening the participation of underrepresented groups in the STEM disciplines

Making it relevant Peter Nentwig,David Waddington,2006 Teaching in context has become an accepted and often welcomed way of teaching science in both primary and secondary schools The conference organised by IPN and the University of York Science Education Group Context based science curricula drew on the experience of over 40 science educators and 10 projects The book is arranged in four parts Part A consists of two papers one on situated learning and the other on implementation of new curricula Part B contains descriptions of five major curricula in different countries why they were introduced how they were developed and implemented and evaluation results Part C gives descriptions of three projects that are of smaller scale and their materials are used as interventions in other more conventional curricula There is also a contribution on some fundamental research where modules of work are written to examine how best to design context based curricula Finally Part D consist of two chapters one summarising some of the findings that came out of the chapters in the three earlier parts and the second looks at the future

Survival Handbook for the New Chemistry Instructor Diane M. Bunce,Cinzia M. Muzzi,2004 This book provides an overview of the issues facing new chemistry faculty in preparation for teaching Serving as a reference to answer specific questions new chemistry faculty encounter this book is comparable to sitting down with a colleague in the department and talking through some ideas or gaining some pointers on how to avoid common pitfalls It is the one single place new chemistry faculty can go to find practical information on how to teach and how to prepare for teaching their first course Chapters are written both by established experts in the field and by new professors within their first couple of years of teaching

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Preparing for Your ACS Examination in General Chemistry - the Official Guide LUCY T. EUBANKS,2018-10-15 This guide is separated into first term and second term general chemistry material Each section contains 8 chapters of material that also aligns to most general chemistry textbooks for a seamless addition to study materials for students Each chapter is designed with an introductory section of the material including common representations and where to find this material in a textbook The second section provides worked examples of typical multiple choice questions including how the correct answer

is determined as well as how the incorrect answers were determined Also included for each study problem is a listing of the corresponding practice questions that use that concept The final section is a series of practice problems to test the concepts collectively The key is provided on a separate page for all study and practice problems

Education ,2001

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