

25. The number of  $\sigma$  bonds in  $\text{N}\equiv\text{N}$  is

- a. 1
- b. 2
- c. 3
- d. 4

26. The elements in an ionic compound are held together by

- a. electrostatic forces of attraction.
- b. van der Waals forces
- c. the spin of paired electrons.
- d. the formation of hybrid orbitals.
- e. an electron pair.

27. In every electrolytic and galvanic (voltaic) cell the anode is that electrode

- a. at which oxidation occurs.
- b. which attracts cations.
- c. at which electrons are supplied to the solution.
- d. at which reduction occurs.

28. Metal X was plated from a solution containing cations of X. The passage of 48.25 C deposited 31 mg of X on the cathode. What is the mass of X (in grams) per mole of electrons?

- a. 47
- b. 62
- c. 93
- d. 186

29. In a galvanic (voltaic) cell in which the reaction is  $\text{Cd} + \text{Cu}^{2+} \rightarrow \text{Cu} + \text{Cd}^{2+}$  and the ions are at unit concentration (activity), the cell potential is.



30. In which reaction will an increase in total pressure at constant temperature favor formation of the products?

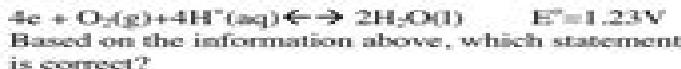
- a.  $\text{CaCO}_3(s) \rightleftharpoons \text{CaO}(s) + \text{CO}_2(g)$
- b.  $\text{H}_2(g) + \text{Cl}_2(g) \rightleftharpoons 2\text{HCl}(g)$
- c.  $2\text{NO}(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}_2(g)$
- d.  $\text{COCl}_2(g) \rightleftharpoons \text{CO}(g) + \text{Cl}_2(g)$

Standard Potentials	$E^\circ$
$\text{Mg} \rightarrow \text{Mg}^{2+} + 2e^-$	2.37 V
$\text{Al} \rightarrow \text{Al}^{3+} + 3e^-$	1.66 V
$\text{Zn} \rightarrow \text{Zn}^{2+} + 2e^-$	0.76 V
$\text{Fe} \rightarrow \text{Fe}^{2+} + 2e^-$	0.44 V
$\text{Cu} \rightarrow \text{Cu}^{2+} + 2e^-$	0.34 V
$\text{Ag} \rightarrow \text{Ag}^{+} + e^-$	-0.80 V

31. Using only the metals Mg, Al, Zn, Fe, Cu and Ag, together with their 1 M salt solutions, a voltaic cell of the highest possible voltage would be constructed using electrodes of these metals.

- a. Mg and Ag
- b. Mg and Fe
- c. Zn and Cu
- d. Al and Ag
- e. Mg and Al

32.  $E = E^\circ - 0.059/n \log Q$  (Nernst equation)  
 $[\text{H}^+] = 1.0 \text{ M}$  initially,  $P_{\text{O}_2} = 1.0 \text{ atm}$



- a.  $n = 1$ , since one mole of oxygen is being considered.
- b. Addition of base should result in an  $E$  value, which is less than 1.23 V.
- c.  $E$  is independent of the pH of the solution.
- d.  $Q = \frac{[\text{H}_2\text{O}]}{[\text{O}_2][\text{H}^+]}$

33. The equilibrium constant for the gaseous reaction  $\text{C} + \text{D} \rightleftharpoons \text{E} + \text{F}$  is 3.0 at 50 °C. In a 2.0 L flask at 50 °C are placed 1.0 mol of C, 1.0 mol of D, 1.0 mol of E, and 3.0 mol of F. Initially, the reaction will

- a. proceed at equal rates in both directions.
- b. proceed more rapidly to form E and F.
- c. proceed more rapidly to form C and D.
- d. not occur in either direction.

Compound	$\Delta G^\circ_f, \text{ kJ/mol}$
$\text{H}_2\text{O}(\text{l})$	-237
$\text{H}_2\text{O}(\text{g})$	-229

34. At 298 K the equilibrium constant for  $\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightleftharpoons \text{H}_2\text{O}(\text{l})$

- a. is larger than the  $K_{\text{eq}}$  for  $\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightleftharpoons \text{H}_2\text{O}(\text{g})$ .
- b. will have a value of 1.0 at equilibrium.
- c. cannot be computed since data on  $\text{O}_2$  and

# Acs Standardized Exam General Chemistry Ii

**Diane M. Bunce,Cinzia M. Muzzi**

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*Guide to Educational Credit by Examination* Douglas R. Whitney, Andrew G. Malizio, 1987 **Chemists' Guide to Effective Teaching** Norbert J. Pienta, Melanie M. Cooper, Thomas J. Greenbowe, 2005 Part of the Prentice Hall Series in Educational Innovation for Chemistry this unique book is a collection of information examples and references on learning theory teaching methods and pedagogical issues related to teaching chemistry to college students In the last several years there has been considerable activity and research in chemical education and the materials in this book integrate the latest developments in chemistry Each chapter is written by a chemist who has some expertise in the specific technique discussed has done some research on the technique and has applied the technique in a chemistry course **Abstracts of Papers** American Chemical Society. Meeting, **Griswold Chem Im 6/E** Nebergall, 1980 **Book of Abstracts** American Chemical Society. Meeting, American Chemical Society, 1987 **Metropolitan Universities**, 2006 **Modeling and Simulation**, 1991

**Preparing for Your ACS Examination in General Chemistry - the Official Guide** LUCY T. EUBANKS, 2018-10-15 This guide is separated into first term and second term general chemistry material Each section contains 8 chapters of material that also aligns to most general chemistry textbooks for a seamless addition to study materials for students Each chapter is designed with an introductory section of the material including common representations and where to find this material in a textbook The second section provides worked examples of typical multiple choice questions including how the correct answer

is determined as well as how the incorrect answers were determined Also included for each study problem is a listing of the corresponding practice questions that use that concept The final section is a series of practice problems to test the concepts collectively The key is provided on a separate page for all study and practice problems     **Journal of Engineering Education** ,2001

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## Table of Contents Acs Standardized Exam General Chemistry Ii

1. Understanding the eBook Acs Standardized Exam General Chemistry Ii
  - The Rise of Digital Reading Acs Standardized Exam General Chemistry Ii
  - Advantages of eBooks Over Traditional Books
2. Identifying Acs Standardized Exam General Chemistry Ii
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Acs Standardized Exam General Chemistry Ii
  - User-Friendly Interface
4. Exploring eBook Recommendations from Acs Standardized Exam General Chemistry Ii
  - Personalized Recommendations
  - Acs Standardized Exam General Chemistry Ii User Reviews and Ratings
  - Acs Standardized Exam General Chemistry Ii and Bestseller Lists
5. Accessing Acs Standardized Exam General Chemistry Ii Free and Paid eBooks
  - Acs Standardized Exam General Chemistry Ii Public Domain eBooks
  - Acs Standardized Exam General Chemistry Ii eBook Subscription Services
  - Acs Standardized Exam General Chemistry Ii Budget-Friendly Options
6. Navigating Acs Standardized Exam General Chemistry Ii eBook Formats
  - ePub, PDF, MOBI, and More
  - Acs Standardized Exam General Chemistry Ii Compatibility with Devices
  - Acs Standardized Exam General Chemistry Ii Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Acs Standardized Exam General Chemistry Ii
  - Highlighting and Note-Taking Acs Standardized Exam General Chemistry Ii
  - Interactive Elements Acs Standardized Exam General Chemistry Ii
8. Staying Engaged with Acs Standardized Exam General Chemistry Ii

- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Acs Standardized Exam General Chemistry Ii
9. Balancing eBooks and Physical Books Acs Standardized Exam General Chemistry Ii
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Acs Standardized Exam General Chemistry Ii
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Acs Standardized Exam General Chemistry Ii
- Setting Reading Goals Acs Standardized Exam General Chemistry Ii
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Acs Standardized Exam General Chemistry Ii
- Fact-Checking eBook Content of Acs Standardized Exam General Chemistry Ii
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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