

# ACIDS, BASES AND SALTS

<b>Definition of an acid</b> An acid is a compound which when dissolved in water produces hydrogen ions as the only positively charged ions	<b>What causes the acidic properties of acids?</b> The hydrogen ions ( $H^+$ ) cause the acidic properties and these are formed in the presence of water.
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<b>Another term to refer to an acid</b> An acid is called a proton donor	<b>Why an acid is also called a proton donor?</b> It's because an acid provides protons or hydrogen ions ( $H^+$ ) to other substances during the reaction.
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<b>Substance to which an acid provides protons</b> BASE	<b>A base is then called</b> PROTON ACCEPTOR	<b>Why the base is called a proton acceptor?</b> Because the base accepts hydrogen ions from acids	<b>Equation for the reaction between acid and base</b> $H_{(aq)} + OH_{(aq)} \longrightarrow H_2O_{(l)}$
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<b>Common laboratory acids</b>		<b>These three common laboratory acids are also called</b> MINERAL ACIDS	<b>Why these acids are also called mineral acids?</b> They are derived from mineral salts ie chlorides for $HCl$ , sulphates for $H_2SO_4$ and nitrates for $HNO_3$
Hydrochloric acid	$HCl$		
Sulphuric acid	$H_2SO_4$		
Nitric acid	$HNO_3$		

<b>Other mineral acids known</b>	<b>Mineral salts from which the acid is derived</b>	<b>Organic acids known</b>	<b>Naturally occurring acids</b>
Sulphurous acid	$H_2SO_3$	Derived from sulphites	CITRIC ACID from lemons
Carbonic acid	$H_2CO_3$	Derived from carbonates	TARTARIC ACID from grapes
Phosphoric acid	$H_3PO_4$	Derived from phosphates	ACETIC ACID from vinegar
Nitrous acid	$HNO_2$	Derived from nitrates	LACTIC ACID from sour milk
		Methanoic acid ( $HCOOH$ )	Hydrochloric acid from digestive juices

<b>Whenever an acid is dissolved in water, it produces</b> HYDROGEN IONS	<b>Term given to the number of hydrogen ions produced by one molecule of an acid</b> BASICITY OF AN ACID	<b>Definition of basicity of an acid</b> BASICITY of an acid is the <u>number of hydrogen ions produced by one molecule of an acid in aqueous solution</u> .
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# Acid Bases Study Guide

**Gerard J. Criner, Gilbert E. D'Alonzo**

## **Acid Bases Study Guide:**

**Acid-Base Equilibria - Quick Chemistry Review Outline and Handout** E Staff, Acid Base Equilibria Quick Review Outline and Handout for All Students Learn and review on the go Use Quick Review Chemistry Notes to help you learn or brush up on the subject quickly You can use the review notes as a reference to understand the subject better and improve your grades Easy to remember facts to help you perform better Perfect study notes for all high school and college students 10 Pages

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