

Acids and bases review sheet Answer key

1. List the differences between an acid and a base. Your description should include such things as whether both are electrolytes, which contains a hydrogen ion and which includes a hydroxide ion, and you should also include two examples of an acid and two examples of a base that are present in your household.

Acid	base
<ul style="list-style-type: none"> <li>Conducts electricity</li> <li>Contains hydrogen ion</li> <li>Proton donor</li> <li>Turns litmus paper red</li> <li>React with most metals</li> <li>Releases hydronium</li> <li>Ex: lemon</li> <li>Vinegar</li> <li>Battery acid (sulfuric acid)</li> </ul>	<ul style="list-style-type: none"> <li>conducts electricity</li> <li>contains hydroxide ion</li> <li>proton acceptor</li> <li>turns litmus paper blue</li> <li>releases hydroxide ion</li> <li>Ex: Windex</li> <li>bleach</li> <li>soap</li> </ul>

2. Listed below are some of the properties of acids and bases. Fill in the blanks with the appropriate word, acids or bases:

- a)  A produce hydrogen ions ( $H^+$ ) in solution
- b)  B have a bitter taste
- c)  B produce hydroxide ions ( $OH^-$ ) ions
- d)  A react with bases to form salts
- e)  A have a sour taste
- f)  B have a slippery, soapy feel
- g)  B react with acids to form salts
- h)  A react with many metals

3. Write an equation for the dissociation of the following acids in water. Follow example.

a)  $\text{HClO}_4 \rightarrow \text{H}^+ + \text{ClO}_4^-$       d)  $\text{H}_2\text{S} \rightarrow 2\text{H}^+ + \text{S}^{2-}$   
 b)  $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$       e)  $\text{HCl} \rightarrow \text{H}^+ + \text{Cl}^-$   
 c)  $\text{HC}_2\text{H}_5\text{O}_2 \rightarrow \text{H}^+ + \text{C}_2\text{H}_5\text{O}_2^-$       f)  $\text{HNO}_3 \rightarrow \text{H}^+ + \text{NO}_3^-$

4. Write an equation for the dissociation (ionization) of the following bases in water: (follow example).

a)  $\text{Mg}(\text{OH})_2 \rightarrow \text{Mg}^{2+} + 2\text{OH}^-$   
b)  $\text{NaOH} \rightarrow \text{Na}^+ + \text{OH}^-$   
c)  $\text{Ba}(\text{OH})_2 \rightarrow \text{Ba}^{2+} + 2(\text{OH}^-)$   
d)  $\text{KOH} \rightarrow \text{K}^+ + \text{OH}^-$   
e)  $\text{Mg}(\text{OH})_2 \rightarrow \text{Mg}^{2+} + 2(\text{OH}^-)$   
f)  $\text{LiOH} \rightarrow \text{Li}^+ + \text{OH}^-$

Q: What type of reaction is a neutralization reaction? Double replacement

B. Write the net equation for neutralization:



# Acid Base Review Answers

**P. W. Bedford**

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