

H. 59. Alpha particles with their high ionizing charge but low penetrating ability can best be shielded by:

a. paper      b. clothing      c. water      d. lead

Alpha  $\alpha$  particles - paper is a sufficient shield

Beta  $\beta$  particles - heavy clothing is sufficient to shield

Gamma rays - lead is required to shield

 = polar

 D = Dipole-Dipole

 = non-polar

 = non-polar

 = non-polar

Mass of Se-72.1	Number of half-lives	Time
40 g	0	0
20 g	1	10 days
10 g	2	20 days
5 g	3	30 days

$\ominus$  = particles go faster  
 $\oplus$  = particles go slower

### Periodic Trends

ATOMIC RADIUS: decreased  $\rightarrow$   
increased  $\downarrow$

METALLIC CHARACTER:

decreases  $\downarrow$   
increases  $\rightarrow$

IONIZATION ( $\ominus$ ) decreases  $\downarrow$   
increases  $\rightarrow$

mass # (# protons + # neutrons)

A 

Z 

• Isotopes = atoms with same # of protons, but different # of neutrons.

Example: Ne-20  $\rightarrow$   $^{20}_{10}$ Ne

Ne-21  $\rightarrow$   $^{21}_{10}$ Ne

Ne-22  $\rightarrow$   $^{22}_{10}$ Ne

Dilution Equation  $M_1 V_1 = M_2 V_2$

• mass % =  $\frac{\text{mass solute}}{\text{mass of solution}} \times 100\%$   
(solute + solvent)

• Molarity =  $\frac{\text{mol solute}}{\text{L solution}}$   
(M)

• molality =  $\frac{\text{mol solute}}{\text{kg solvent}}$

• # protons - # electrons = Charge #

• # neutrons = A (mass #) - Z (atomic #)

• if no charge, the # protons + # electrons = same

• if + (positive) Charge (cation) = # electrons is less than # protons

• if - (negative) Charge (anion) = # electrons is greater than # protons

• Protons + Charge  $\rightarrow$  they give identity to the element.

# Acs General Chemistry Exam Review

**Christian Drosten**

## **Acs General Chemistry Exam Review:**

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