

# Inorganic Chemistry ACS: Study Guide Exam|334 Questions with Accurate Answers

Ionization energy - -energy required to remove the least tightly bound electron from a neutral atom in the gas phase

- periodic trend of ionization energy - -highest at top right-smaller electron=harder to remove
- Why is a half filled subshell so stable? - -it serves to maximize the stabilizing interactions while minimizing the destabilizing interactions among electrons
- exchange interaction - -pic, stabilizing, result of electrons pairing in degenerate orbitals with parallel spin
- pairing energy - -destabilizing, coulomb interaction, pic, energy of electron-electron repulsion in a filled orbital
- Is it easier to ionize a high energy or low energy electrons - -high energy electron-already contains more energy so it requires less energy input
- What happens when a 3d series metal is ionized? - -the first electron to be ionized will come from the 4s orbital, the other's electron will enter the d orbital ( $4s03dn+1$ )
- lanthanide contraction - -reduction in atomic radius following the lanthanide series, contrary to the overall trend observed for the periodic table
- lanthanides - -elements 57-71, first appearance of f orbitals, f orbitals are poor at shielding so any electrons added will have a higher  $Z_{eff}$ , shrinking the radius
- Slater's rules - -tell us what the effective nuclear charge will be,  $Z_{eff}=Z-\sigma$ , Z is the atomic number,  $\sigma$ =sum of the number of electrons in a given shell multiplied by a weighting coefficient (page 1)
- Shielding - -the reduction in charge attraction between the nucleus and electrons due to electrons between the nucleus and the electron in question, it is considered to be between if it has a lower energy

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*Instructor* Diane M. Bunce, Cinzia M. Muzzi, 2004 This book provides an overview of the issues facing new chemistry faculty in preparation for teaching. Serving as a reference to answer specific questions new chemistry faculty encounter, this book is comparable to sitting down with a colleague in the department and talking through some ideas or gaining some pointers on how to avoid common pitfalls. It is the one single place new chemistry faculty can go to find practical information on how to teach and how to prepare for teaching their first course. Chapters are written both by established experts in the field and by new professors within their first couple of years of teaching.

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