

- Structure, Hybridization, Resonance, Aromaticity
 - # of bonds- 1=sp (e.g. 3 bonds = sp^2)
 - Aromaticity= $4n+2$ electrons
 - Strained ring causes greater heat of combustion
- Acids and Bases
 - Lewis acid= accepts electron pair
 - Higher percentage of s character means more acidic hydrogen
- Stereoisomerism
 - Constitutional Isomers
 - Isomers that have different atom connectivity
 - Stereoisomers
 - Same connectivity but different arrangement in space
 - Enantiomers
 - Stereoisomers that are nonsuperimposable mirror reflections
 - Diastereomers
 - Stereoisomers that are not mirror reflections of each other
 - Stereocenter= sp^3 carbon bonded to four different substituents
 - Assigning R/S configuration
 - If lowest group is pointing away, determine normally
 - If lowest group is not pointing away, exchange with group pointing away and determine reverse
- Nucleophilic Substitutions and Eliminations
 - | Alkyl Halide | Strong Base | Weak Base | Poor Nucleophile |
|--------------|-------------|-----------------------------------|------------------------------------|
| 1 | $Sn2$ | $Sn2$ | No rxn. |
| 2 | Mostly $E2$ | $Sn2$ | $Sn1/E1$ in polar, protic solvents |
| 3 | $E2$ | $Sn1/E1$ in polar, protic solvent | $Sn1/E1$ in polar, protic solvents |
 - Relative strength of nucleophiles
 - Negatively charged nucleophiles react faster in substitution reactions than neutral nucleophiles
 - Nucleophilic strength is in the same order as base strength when the nucleophilic atoms are in the same period of the periodic table
 - Nucleophilic strength increases with increasing atom size when the nucleophilic atoms are in the same column of the periodic table
 - Leaving group ability
 - Increases with decreasing base strength
 - Base strength and structure

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