

# Acids and Bases

MERCK

Acid				Base			
Strength of Acid	Ka	Name	Formula	Formula	Name	Strength of Base	
<b>Strongest</b>	Large	Perchloric acid	HClO <sub>4</sub>	ClO <sub>4</sub> <sup>-</sup>	Perchlorate ion	<b>Weakest</b>	
	3.2×10 <sup>8</sup>	Hydroiodic acid		I <sup>-</sup>	Iodide		
	1.0×10 <sup>9</sup>	Hydrobromic acid	HBr	Br <sup>-</sup>	Bromide		
	1.3×10 <sup>4</sup>	Hydrochloric acid	HCl	Cl <sup>-</sup>	Chloride		
	1.0×10 <sup>3</sup>	Sulfuric acid	H <sub>2</sub> SO <sub>4</sub>	HSO <sub>4</sub> <sup>-</sup>	Hydrogen sulfate ion		
	2.4×10 <sup>1</sup>	Nitric acid	HNO <sub>3</sub>	NO <sub>3</sub> <sup>-</sup>	Nitrate ion		
	.....	Hydronium ion	H <sub>3</sub> O <sup>+</sup>	H <sub>2</sub> O	Water		
	5.4×10 <sup>-2</sup>	Oxalic acid	HO <sub>2</sub> C <sub>2</sub> O <sub>2</sub> H	HO <sub>2</sub> C <sub>2</sub> O <sub>2</sub> <sup>-</sup>	Hydrogen oxalate ion		
	1.3×10 <sup>-2</sup>	Sulfurous acid	H <sub>2</sub> SO <sub>3</sub>	HSO <sub>3</sub> <sup>-</sup>	Hydrogen sulfite ion		
	1.0×10 <sup>-2</sup>	Hydrogen sulfate ion	HSO <sub>4</sub> <sup>-</sup>	SO <sub>4</sub> <sup>2-</sup>	Sulfate ion		
	7.1×10 <sup>-3</sup>	Phosphoric acid	H <sub>3</sub> PO <sub>4</sub>	H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>	Dihydrogen phosphate ion		
	7.2×10 <sup>-4</sup>	Nitrous acid	HNO <sub>2</sub>	NO <sub>2</sub> <sup>-</sup>	Nitrite ion		
	6.6×10 <sup>-4</sup>	Hydrofluoric acid	HF	F <sup>-</sup>	Fluoride ion		
	1.8×10 <sup>-4</sup>	Methanoic acid	HCO <sub>2</sub> H	HCO <sub>2</sub> <sup>-</sup>	Methanoate ion		
	6.3×10 <sup>-5</sup>	Benzoic acid	C <sub>6</sub> H <sub>5</sub> COOH	C <sub>6</sub> H <sub>5</sub> COO <sup>-</sup>	Benzoate ion		
	5.4×10 <sup>-5</sup>	Hydrogen oxalate ion	HO <sub>2</sub> C <sub>2</sub> O <sub>2</sub> <sup>-</sup>	O <sub>2</sub> C <sub>2</sub> O <sub>2</sub> <sup>2-</sup>	Oxalate ion		
	1.8×10 <sup>-5</sup>	Ethanoic acid	CH <sub>3</sub> COOH	CH <sub>3</sub> COO <sup>-</sup>	Ethanoate (acetate) ion		
	4.4×10 <sup>-7</sup>	Carbonic acid	H <sub>2</sub> CO <sub>3</sub>	HCO <sub>3</sub> <sup>-</sup>	Hydrogen carbonate ion		
	1.1×10 <sup>-7</sup>	Hydrosulfuric acid	H <sub>2</sub> S	HS <sup>-</sup>	Hydrogen sulfide ion		
	6.3×10 <sup>-8</sup>	Dihydrogen phosphate ion	H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>	HPO <sub>4</sub> <sup>2-</sup>	Hydrogen phosphate ion		
	6.2×10 <sup>-8</sup>	Hydrogen sulfite ion	HS <sup>-</sup>	SO <sub>3</sub> <sup>2-</sup>	Sulfite ion		
	2.9×10 <sup>-8</sup>	Hypochlorous acid	HClO	ClO <sup>-</sup>	Hypochlorite ion		
	6.2×10 <sup>-10</sup>	Hydrocyanic acid	HCN	CN <sup>-</sup>	Cyanide ion		
	5.8×10 <sup>-10</sup>	Ammonium ion	NH <sub>4</sub> <sup>+</sup>	NH <sub>3</sub>	Ammonia		
	5.8×10 <sup>-10</sup>	Boric acid	H <sub>3</sub> BO <sub>3</sub>	H <sub>2</sub> BO <sub>3</sub> <sup>-</sup>	Dihydrogen borate		
	4.7×10 <sup>-11</sup>	Hydrogen carbonate ion	HCO <sub>3</sub> <sup>-</sup>	CO <sub>3</sub> <sup>2-</sup>	Carbonate ion		
	4.2×10 <sup>-13</sup>	Hydrogen phosphate ion	HPO <sub>4</sub> <sup>2-</sup>	PO <sub>4</sub> <sup>3-</sup>	Phosphate ion		
	1.8×10 <sup>-13</sup>	Dihydrogen borate ion	H <sub>2</sub> BO <sub>3</sub> <sup>-</sup>	HBO <sub>3</sub> <sup>2-</sup>	Hydrogen borate ion		
	1.3×10 <sup>-13</sup>	Hydrogen sulfide ion	HS <sup>-</sup>	S <sup>2-</sup>	Sulfide ion		
	1.6×10 <sup>-14</sup>	Hydrogen borate ion	HBO <sub>3</sub> <sup>2-</sup>	BO <sub>3</sub> <sup>3-</sup>	Borate ion		
<b>Weakest</b>	.....	water	H <sub>2</sub> O	OH <sup>-</sup>	Hydroxide	<b>Strongest</b>	



# Acids And Bases Guide

**Angela Randels-Thorp, David Liss**



## Acids And Bases Guide:

**Guide to Acids & Bases Kit** Centre for Adult Education, Shell South Africa (Pty) Limited, 1994 *Pocket Guide to Fluid, Electrolyte, and Acid-Base Balance - E-Book* Ursula Heitz, Mima M. Horne, 2004-10-20 Compares different types of parenteral and enteral feedings along with nursing implications Contains expanded and updated relevant IV therapy and pharmacology Features updated content throughout with current literature and research findings such as most current blood pressure guidelines from the U S Department of Health and Human Services *Handbook of Acid-Base Indicators* R. W. Sabnis, 2007-10-04 While acid base indicators continue to find new applications in an ever widening range of scientific disciplines there is no current book that focuses entirely on the subject nor one that brings together the relevant advances that have evolved over the last three decades The Handbook of Acid Base Indicators compiles the most up to date c  
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### Table of Contents **Acids And Bases Guide**

1. Understanding the eBook **Acids And Bases Guide**
  - The Rise of Digital Reading **Acids And Bases Guide**
  - Advantages of eBooks Over Traditional Books
2. Identifying **Acids And Bases Guide**
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an **Acids And Bases Guide**
  - User-Friendly Interface
4. Exploring eBook Recommendations from **Acids And Bases Guide**
  - Personalized Recommendations
  - **Acids And Bases Guide** User Reviews and Ratings
  - **Acids And Bases Guide** and Bestseller Lists

5. Accessing Acids And Bases Guide Free and Paid eBooks
  - Acids And Bases Guide Public Domain eBooks
  - Acids And Bases Guide eBook Subscription Services
  - Acids And Bases Guide Budget-Friendly Options
6. Navigating Acids And Bases Guide eBook Formats
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7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Acids And Bases Guide
  - Highlighting and Note-Taking Acids And Bases Guide
  - Interactive Elements Acids And Bases Guide
8. Staying Engaged with Acids And Bases Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Acids And Bases Guide
9. Balancing eBooks and Physical Books Acids And Bases Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Acids And Bases Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Acids And Bases Guide
  - Setting Reading Goals Acids And Bases Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Acids And Bases Guide
  - Fact-Checking eBook Content of Acids And Bases Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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