

# Acids and Bases

MERCK

Acid			Base			
Strength of Acid	Ka	Name	Formula	Formula	Name	Strength of Base
<b>Strongest</b>	Large	Perchloric acid	HClO <sub>4</sub>	ClO <sub>4</sub> <sup>-</sup>	Perchlorate ion	
	$3.2 \times 10^{-1}$	Hydroiodic acid	HBr	I <sup>-</sup>	Iodide	
	$1.0 \times 10^{-1}$	Hydrobromic acid	HCl	Br <sup>-</sup>	Bromide	
	$1.3 \times 10^{-1}$	Hydrochloric acid	H <sub>2</sub> SO <sub>4</sub>	Cl <sup>-</sup>	Chloride	
	$1.0 \times 10^{-2}$	Sulfuric acid	HNO <sub>3</sub>	HSO <sub>4</sub> <sup>-</sup>	Hydrogen sulfate ion	
	$2.4 \times 10^{-3}$	Nitric acid	H <sub>3</sub> O <sup>+</sup>	NO <sub>3</sub> <sup>-</sup>	Nitrate ion	
		Hydroxium ion	H <sub>3</sub> O <sup>+</sup>	H <sub>2</sub> O	Water	
	$5.4 \times 10^{-4}$	oxalic acid	HO <sub>2</sub> C <sub>2</sub> O <sub>4</sub> H	HO <sub>2</sub> C <sub>2</sub> O <sub>4</sub> <sup>2-</sup>	Hydrogen oxalate ion	
	$1.3 \times 10^{-4}$	Sulfurous acid	H <sub>2</sub> SO <sub>3</sub>	HSO <sub>3</sub> <sup>-</sup>	Hydrogen sulfite ion	
	$1.0 \times 10^{-4}$	Hydrogen sulfate ion	HSO <sub>4</sub> <sup>-</sup>	SO <sub>4</sub> <sup>2-</sup>	Sulfate ion	
	$7.1 \times 10^{-5}$	Phosphoric acid	H <sub>3</sub> PO <sub>4</sub>	H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>	Dihydrogen phosphate ion	
	$7.2 \times 10^{-5}$	Nitrous acid	HNO <sub>2</sub>	NO <sub>2</sub> <sup>-</sup>	Nitrite ion	
	$6.6 \times 10^{-6}$	Hydrofluoric acid	HF	F <sup>-</sup>	Fluoride ion	
	$1.8 \times 10^{-6}$	Methanolic acid	HCO <sub>2</sub> H	HCO <sub>3</sub> <sup>-</sup>	Methanoate ion	
	$6.3 \times 10^{-7}$	Benzoic acid	C <sub>6</sub> H <sub>5</sub> COOH	C <sub>6</sub> H <sub>5</sub> COO <sup>-</sup>	Benzoate ion	
	$5.4 \times 10^{-7}$	Hydrogen oxalate ion	HO <sub>2</sub> C <sub>2</sub> O <sub>4</sub> <sup>2-</sup>	O <sub>2</sub> C <sub>2</sub> O <sub>4</sub> <sup>2-</sup>	Oxalate ion	
	$1.8 \times 10^{-7}$	Ethanoic acid	CH <sub>3</sub> COOH	CH <sub>3</sub> COO <sup>-</sup>	Ethanoate (acetate) ion	
	$4.4 \times 10^{-7}$	Carbonic acid	H <sub>2</sub> CO <sub>3</sub>	HCO <sub>3</sub> <sup>-</sup>	Hydrogen carbonate ion	
	$1.1 \times 10^{-7}$	Hydrosulfuric acid	H <sub>2</sub> S	HS <sup>-</sup>	Hydrogen sulfide ion	
	$6.3 \times 10^{-8}$	Dihydrogen phosphate ion	H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>	HPO <sub>4</sub> <sup>2-</sup>	Hydrogen phosphate ion	
	$6.2 \times 10^{-8}$	Hydrogen sulfite ion	HS <sup>-</sup>	SO <sub>3</sub> <sup>2-</sup>	Sulfite ion	
	$2.9 \times 10^{-8}$	Hypochlorous acid	HClO	ClO <sup>-</sup>	Hypochlorite ion	
	$6.2 \times 10^{-10}$	Hydrocyanic acid	HCN	CN <sup>-</sup>	Cyanide ion	
	$5.8 \times 10^{-10}$	Ammonium ion	NH <sub>4</sub> <sup>+</sup>	NH <sub>3</sub>	Ammonia	
	$5.8 \times 10^{-10}$	Boric acid	H <sub>3</sub> BO <sub>3</sub>	H <sub>2</sub> BO <sub>3</sub> <sup>-</sup>	Dihydrogen borate	
	$4.7 \times 10^{-11}$	Hydrogen carbonate ion	HCO <sub>3</sub> <sup>-</sup>	CO <sub>3</sub> <sup>2-</sup>	Carbonate ion	
	$4.2 \times 10^{-11}$	Hydrogen phosphate ion	HPO <sub>4</sub> <sup>2-</sup>	PO <sub>4</sub> <sup>3-</sup>	Phosphate ion	
	$1.8 \times 10^{-12}$	Dihydrogen borate ion	H <sub>2</sub> BO <sub>3</sub> <sup>-</sup>	HBO <sub>3</sub> <sup>2-</sup>	Hydrogen borate ion	
	$1.3 \times 10^{-12}$	Hydrogen sulfide ion	HS <sup>-</sup>	S <sup>2-</sup>	Sulfide ion	
	$1.6 \times 10^{-12}$	Hydrogen borate ion	HBO <sub>3</sub> <sup>2-</sup>	BO <sub>3</sub> <sup>3-</sup>	Borate ion	
<b>Weakest</b>	-----	water	H <sub>2</sub> O	OH <sup>-</sup>	Hydroxide	<b>Strongest</b>



# Acids And Bases Guide

**Angela Randels-Thorp,David Liss**

## Acids And Bases Guide:

**Guide to Acids & Bases Kit** Centre for Adult Education,Shell South Africa (Pty) Limited,1994      *Pocket Guide to Fluid, Electrolyte, and Acid-Base Balance - E-Book* Ursula Heitz, Mima M. Horne,2004-10-20 Compares different types of parenteral and enteral feedings along with nursing implications Contains expanded and updated relevant IV therapy and pharmacology Features updated content throughout with current literature and research findings such as most current blood pressure guidelines from the U S Department of Health and Human Services      *Handbook of Acid-Base Indicators* R. W. Sabnis,2007-10-04 While acid base indicators continue to find new applications in an ever widening range of scientific disciplines there is no current book that focuses entirely on the subject nor one that brings together the relevant advances that have evolved over the last three decades The Handbook of Acid Base Indicators compiles the most up to date c

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Randels-Thorp,David Liss,2016-12-19 Acid Base and Electrolyte Handbook for Veterinary Technicians provides an easy to understand yet comprehensive approach to acid base and electrolyte balance Covers the physiology of fluids and their effect on acid base and electrolyte balance Offers detailed information on managing acid base and electrolyte derangements in disease Includes access to a companion website with case studies and multiple choice questions      **Acids and Bases**

**Teacher's Guide** Tina Youngman,2017 This workbook is designed to give students an understanding of Acids and Bases and will lead students into Level 1 NCEA It also includes material on writing equations and formulae It focuses on aspects of Levels 3 4 and 5 of the New Zealand Science Curriculum It gives students a base to build onto with reference to the properties of acids and bases the pH scale indicators word and basic symbol equations neutralisation and how humans use acids and bases The literacy tasks in this workbook encourage the students to read for meaning scan material select

appropriate information and understand key words The inclusion of mini research projects prepares students for what is required in the Senior Sciences The numeracy tasks involve number skills rounding using basic formulae graphing data and reading tables accurately These are all vital skills in the Sciences and beyond Preface **Fundamentals of Organic Chemistry, Textbook, Study Guide and Solutions Manual** T. W. Graham Solomons,1996-08-15 A realistic approach to the study of mechanisms The book addresses real functional group chemistry with an emphasis on the biological environmental and medical applications of organic chemistry **Smith's Practical Dyer's Guide** David Smith,1881

**Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition** Susan McMurry,2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Price winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines

**Handbook of Nutritional Disorders** Jahangir Moini,Katia Ferdowsi,2024-12-10 Handbook of Nutritional Disorders is a comprehensive handbook covering topics in nutrition malnutrition and the clinical disorders associated with nutrition from deficiency to toxicity It includes information on disorders related to carbohydrate lipid and protein metabolism as well as vitamin and mineral abnormalities The book details various types of supplements feeding methods and therapies for many specific patients It aims to educate readers on ways to prevent disorders resulting from all types of malnutrition and their potentially severe complications Features Strong focus on diabetes featuring information on various forms of the disease and treatment information Detailed discussion of lipids and related disorders cardiovascular disease is the number one cause of death informs users on prevention and treatment of hypertension myocardial infarction and stroke Contains information on selective nutritional disorders including obesity dehydration imbalances malabsorption alcoholism neuropsychiatric disorders eating disorders cancer and pollutant poisonings Written for researchers academia and students in healthcare and nutrition fields this book educates individuals on prevention of disorders resulting from all types of malnutrition and their potentially severe complications **Manual of Chemistry** William Simon,1898 **Laboratory manual** William McPherson,1913

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