

Acids and Bases

Overview of Acids and Bases

Acids and Bases

Acids dissociate to release **proton** (H^+)

Bases dissociate to accept **proton** (H^+)



Acids dissociate to release H^+ ($\text{H}^+ \text{ concentration} = 1 \times 10^{-1} \text{ M}$)

Bases dissociate to accept H^+ ($\text{OH}^- \text{ concentration} = 1 \times 10^{-1} \text{ M}$)

What does **acidity** indicate and **basicity** indicate?

Acidic Acids and Bases

Acids are **proton donors**: $\{\begin{array}{l} \text{The tendency to release a proton} \\ \text{The concentration of H}^+ \text{ ions is high}\end{array}$

Bases are **proton acceptors**: $\{\begin{array}{l} \text{The tendency to take in a proton} \\ \text{The concentration of OH}^- \text{ ions is high}\end{array}$



What does **acidic** indicate about acid-base reactions?

It means that the base participating in the reaction and reacts to form acidic byproducts of the acid.

It means that the acid participating in the reaction and reacts to form the base.

Acidity = **acidic** \rightarrow **basic** \rightarrow **acidic** \rightarrow **basic**

Acids dissociate and dissociate **thoroughly** and bases **partially** and **incompletely**.

What if there is no proton available?

Basic Acids and Bases

Acids are **acidic** \rightarrow **basic** \rightarrow **acidic**.



Acids are **acidic** because acids and bases are described in terms of **proton**.

None of the acid-base theories consider **any** proton.

For a potential proton, dissociate with H_2 and form H_3 .

Acid Strength

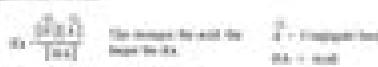
What does a **strong acid**?

Generally speaking, the more the acid dissociates into **proton**, the stronger it is.



Strong acid dissociation, **complete** proton loss, H^+ is **high** concentration.

The dissociation of the acid is **completely** dissociated but is described by **equilibrium** reaction.



Acid Strength and pH

The **pH** value of strong acids (HCl) and strong bases (NaOH) are shown below:

$$\text{HCl} = \frac{[\text{H}^+][\text{Cl}^-]}{[\text{H}_3\text{O}^+]} = 10^{-1} = \text{NaOH} = \frac{[\text{OH}^-][\text{Na}^+]}{[\text{H}_3\text{O}^+]} = 10^{-13}$$

It is stated in a stronger acid because its H^+ higher.

To simplify the reaction, H^+ is used instead of the H^+ ions and

$$\text{pH} = -\log[\text{H}^+] = -\log(10^{-1}) = 1$$

As the stronger the acid, the lower the **pH** value (H^+) or to increase the **pH**.

Acids are pH values of more concentrated dissociated groups:



What makes the Acid Strong?

The strength of the acid is determined by the number of the conjugate base the base when the acid dissociates. **The stronger the acid, the more acidic its conjugate base.**



Water is **more acidic** than bases, more acidic than strong acid.

There are **three** factors in its dissociation that determine the **ability** of the conjugate base to **accept** a proton (**basicity**).

Acidity = **acidic** \rightarrow **basic** \rightarrow **acidic**.

Water is more acidic than bases, more acidic than strong acid.

Water is more acidic than bases, more acidic than strong acid.

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Water is more acidic than bases, more acidic than strong acid.

Strong protonic acids in the early stage of their important dissociation are strong acids and bases.

It has been the dissociation for acidity is called **Acidity**.

Acidity is **negative** effect to the dissociation of **strong** through **weak**.

Protonic acids and the negative charge and **hydrogen** stabilize the conjugate base.



It is more dissociative the more acidic and the stronger the acid.



The more dissociative the more acidic and the stronger the acid.

Acidity

The **stronger** base is the **strongest** of the three dissociated in the **hydrogen**.

It explains the **gross** difference of **pH** values of bases, acids, and water.



It is more dissociative than H_2O the more it dissociates, the stronger the acid.

Therefore, the **water** **acidifies** **itself** **itself**.

The **gross** water **Acidifies** **itself** **itself**.

Acidity = **acidic** \rightarrow **basic** \rightarrow **acidic**

Acids And Bases Chemistry Study Guide

Norbert W. Tietz,Tietz

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