

Acids and Bases

Overview of Acids and Bases

Indicators: Acids and Bases

Acids: Donates its proton to base (proton donor)

Base: Donates its lone pair to form a bond (LBP)

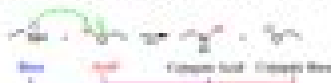


Remember: (1) only H atoms, (2) only OH is a base

What about organic indicators and how organic indicators?

Acids and Bases

Acids are proton donors. (The electrons are being to move)
Bases are proton acceptors. (The electrons are being to move)



General acid-base reactions of acid-base reactions:

to find the strong base, the base gets protonated, change of the base

and usually for weak acids, hydrogens of the acid.

to find the level of the hydrogens and move the electrons

to the other atom

Remember: -OH is a weak base, -OH is a weak acid

1. Acid: Donates its proton to base (proton donor)

2. Base: Accepts its proton to base (proton acceptor)

3. Conjugate acid: The acid that is left after the base has accepted its proton

4. Conjugate base: The base that is left after the acid has donated its proton

5. Strong acid: An acid that completely dissociates in water

6. Weak acid: An acid that partially dissociates in water

7. Strong base: A base that completely dissociates in water

8. Weak base: A base that partially dissociates in water

9. Amphoteric: A substance that can act as both an acid and a base

10. Buffer: A solution that resists changes in pH

11. Indicator: A substance that changes color at a certain pH

12. Titration: A process used to determine the concentration of an acid or base

13. Equivalence point: The point at which the acid and base have reacted in a 1:1 ratio

14. End point: The point at which the indicator changes color

15. pH: A measure of the acidity or basicity of a solution

16. pKa: A measure of the acidity of a weak acid

17. pKb: A measure of the basicity of a weak base

18. Ka: The acid dissociation constant

19. Kb: The base dissociation constant

20. Kw: The ion-product constant for water

21. Le Chatelier's principle: A principle that describes how a system at equilibrium responds to changes in conditions

22. Common ion effect: A phenomenon that occurs when a common ion is added to a solution at equilibrium

23. Buffer capacity: The ability of a buffer to resist changes in pH

24. Henderson-Hasselbalch equation: An equation that relates the pH of a buffer to the pKa of the weak acid and the concentrations of the weak acid and its conjugate base

25. Polyprotic acid: An acid that can donate more than one proton

26. Amphiprotic: A substance that can act as both an acid and a base

27. Zwitterion: A molecule that has both positive and negative charges

28. Isoelectric point: The pH at which a molecule has no net charge

29. Titration curve: A graph of pH versus volume of titrant added

30. Indicator range: The pH range over which an indicator changes color

31. Buffer range: The pH range over which a buffer is effective

32. Acid-base reaction: A reaction in which an acid and a base react to form a conjugate acid and a conjugate base

33. Proton transfer: The movement of a proton from one molecule to another

34. Electron transfer: The movement of an electron from one molecule to another

35. Redox reaction: A reaction in which both oxidation and reduction occur

36. Oxidation: The loss of electrons or an increase in oxidation state

37. Reduction: The gain of electrons or a decrease in oxidation state

38. Oxidizing agent: A substance that causes another substance to be oxidized

39. Reducing agent: A substance that causes another substance to be reduced

40. Half-reaction: A reaction that shows either oxidation or reduction

41. Standard reduction potential: A measure of the tendency of a chemical species to be reduced

42. Standard oxidation potential: A measure of the tendency of a chemical species to be oxidized

43. Standard cell potential: The potential difference between two half-cells in a galvanic cell

44. Nernst equation: An equation that relates the cell potential to the standard cell potential and the concentrations of the reactants and products

45. Gibbs free energy: A thermodynamic potential that can be used to determine the spontaneity of a process

46. Enthalpy: A thermodynamic quantity representing the total energy of a system

47. Entropy: A thermodynamic quantity representing the degree of disorder or randomness of a system

48. Free energy: The energy available to do work at constant temperature and pressure

49. Equilibrium constant: A constant that describes the ratio of the concentrations of the products to the reactants at equilibrium

50. Reaction quotient: A quantity that indicates the relative amounts of the products and reactants present during a reaction at a particular point in time

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Acids And Bases Chemistry Study Guide

Norbert W. Tietz, Tietz



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Study Guide James E. Brady, John R. Holum, 1993-02-11 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping students develop skills in analytical thinking and problem solving Students are taught in a variety of ways to think of skills as tools that can be used to solve complex problems Several aids are included to help focus and inspire student interest frequent reference to common chemicals in commercial products numerous photographs of reactions in chapter practice exercises following worked examples

Clinical Chemistry Study Guide Norbert W. Tietz, Tietz, 1987

Study Guide for General Chemistry and College Chemistry, Eighth Editions by Holtzclaw and Robinson Norman E. Griswold, Henry Fuller Holtzclaw, 1988

MCAT Complete; Medical College Admission Test Comprehensive Study Guide Physical Sciences; Biological Sciences; Verbal Reasoning James Adams, Maryn Munley, Nathan Eisenbeis, Caroline Miller, Jesse Shetler, 2021-10-12 The MCAT The Medical College Admission Test MCAT is a standardized multiple choice examination designed to assess the examinee's problem solving critical thinking and knowledge of science concepts and principles prerequisite to the study of medicine Scores are reported in Physical Sciences Verbal Reasoning and Biological Sciences Study guide covers AAMC Association of American Medical Colleges content Physical Sciences Biological Sciences Verbal Reasoning Mathematics Concepts The Cell Chromosomes Reproduction Implantation Microorganisms Biochemistry Human Physiology The Heart The Lymphatic System GI Tract Musculoskeletal System Kidney Hormones Nerves Skin Genetics Populations and Evolution Elements Hund's Rule and Radiation The Periodic Table Covalent Bonds Molecular Shapes General Chemistry Kinetic Molecular Theory Phase Change Solutions Oxidation Numbers Entropy Acids and Bases Galvanic and Electrolytic Cells Carbon Stereochemistry Alkanes and Alkenes Hydrogen Bonding Alcohols Phenols Aldehydes and Ketones in Carboxylic Acids Ether Ammonia Amino Acids Carbohydrates Spectroscopy Separation and Distillation Vectors and Simple Motion Forces Circular and Projectile Motion Statics Center of Gravity Work and Energy Power and Momentum Stress and Strain Elasticity and Density Hydrostatic Pressure Fluids in Motion Electricity and Magnetism Plates Capacitors Voltage Batteries Resistors Magnetic Fields Waves and Periodic Motion Sound Waves Doppler Effect Simple Harmonic Motion Optics Mirrors 205 pages

Fundamentals of Organic Chemistry, Textbook, Study Guide and Solutions Manual T. W. Graham Solomons, 1996-08-15 A realistic approach to the study of mechanisms The book addresses real functional group chemistry with an emphasis on the biological environmental and medical applications of organic chemistry

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