

## Acids, Bases and Salts

- THE THEORY of ACIDS and ALKALIS and a few technical terms:
  - Acids are substances that form hydrogen ions ( $H^{+}$ ) when dissolved in water e.g. hydrochloric acid HCl gives  $H^{+}$  and  $Cl^{-}$  ions, sulphuric acid  $H_2SO_4$  gives  $2H^{+}$  and  $SO_4^{2-}$  ions and nitric acid  $HNO_3$  gives  $H^{+}$  and  $NO_3^{-}$  ions.
  - Alkalies are substances that form hydroxide ions ( $OH^{-}$ ) in water e.g. sodium hydroxide NaOH gives  $Na^{+}$  and  $OH^{-}$  ions, calcium hydroxide  $Ca(OH)_2$  gives  $Ca^{2+}$  and  $2OH^{-}$  ions. Note: an alkali is a base soluble in water.
  - In water, there are trace quantities of  $H^{+}$  and  $OH^{-}$  ions BUT they are of equal concentration and so water is neutral.
  - In acid solutions there are more  $H^{+}$  ions than  $OH^{-}$  ions.
  - In alkaline solution there are more  $OH^{-}$  ions than  $H^{+}$  ions.
  - Acids dissociate to different extents in aqueous solution. Acids that dissociate to a large extent are strong electrolytes and strong acids. In contrast, acids that dissociate only to a small extent are weak acids and weak electrolytes.

In a similar manner, bases can be strong or weak depending on the extent to which they dissociate and produce  $OH^{-}$  ions in solution. Most metal hydroxides are strong electrolytes and strong bases. Ammonia,  $NH_3$ , is a weak electrolyte and weak base.



- BASES e.g. oxides and hydroxides are substances that react and neutralise acids to form salts and water. Bases which are soluble in water are called alkalis.

### Acids

Some common acids are listed below:

| Name                    | Formula     | Strong/Weak | Where is it found?                 |
|-------------------------|-------------|-------------|------------------------------------|
| Hydrochloric acid       | HCl         | Strong      | The stomach, in the lab.           |
| Sulphuric acid          | $H_2SO_4$   | Strong      | Acid rain, car batteries, the lab. |
| Nitric acid             | $HNO_3$     | Strong      | Acid rain, in the lab.             |
| Ethanoic (acetic) acid  | $CH_3COOH$  | Weak        | Vinegar                            |
| Methanoic (formic) acid | $HCOOH$     | Weak        | Ant & nettle stings, descalers     |
| Citric Acid             | $C_6H_8O_7$ | Weak        | Citrus fruits                      |

- Acids taste sour (e.g. vinegar, lemon juice).
- Acids are harmful to living cells.
- Aqueous solutions of all acids contain hydrogen ions,  $H^{+}$ .

# Acids Bases And Salts Study Guide

**H. Stephen Stoker,Danny V. White**

## Acids Bases And Salts Study Guide:

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