

Acids, Bases and Salts

- **THE THEORY of ACIDS and ALKALIS** and a few technical terms:
 - **Acids** are substances that **form hydrogen ions ($\text{H}^+_{(\text{aq})}$)** when dissolved in water eg hydrochloric acid HCl gives $\text{H}^+_{(\text{aq})}$ and $\text{Cl}^-_{(\text{aq})}$ ions, sulphuric acid H_2SO_4 gives $2\text{H}^+_{(\text{aq})}$ and SO_4^{2-} ions and nitric acid HNO_3 gives $\text{H}^+_{(\text{aq})}$ and $\text{NO}_3^-_{(\text{aq})}$ ions.
 - **Alkalis** are substances that **form hydroxide ions ($\text{OH}^-_{(\text{aq})}$)** in water eg sodium hydroxide NaOH gives $\text{Na}^+_{(\text{aq})}$ and $\text{OH}^-_{(\text{aq})}$ ions, calcium hydroxide $\text{Ca}(\text{OH})_2$ gives $\text{Ca}^{2+}_{(\text{aq})}$ and $2\text{OH}^-_{(\text{aq})}$ ions. Note: **an alkali is a base soluble in water.**
 - In **water**, there are trace quantities of H^+ and OH^- ions BUT they are of equal concentration and so water is neutral.
 - In **acid solutions** there are **more H^+ ions than OH^- ions.**
 - In **alkaline solution** there are **more OH^- ions than H^+ ions.**
 - Acids dissociate to different extents in aqueous solution. Acids that dissociate to a large extent are strong electrolytes and **strong acids**. In contrast, acids that dissociate only to a small extent are **weak acids** and weak electrolytes.

In a similar manner, bases can be strong or weak depending on the extent to which they dissociate and produce OH^- ions in solution. Most metal hydroxides are strong electrolytes and **strong bases**. Ammonia, NH_3 , is a weak electrolyte and **weak base**.



- **BASES** eg oxides and hydroxides are substances that **react and neutralise acids to form salts and water**. **Bases which are soluble in water are called alkalis.**

Acids

Some common acids are listed below:

| Name | Formula | Strong/Weak | Where is it found? |
|-------------------------|----------------------------------|-------------|------------------------------------|
| Hydrochloric acid | HCl | Strong | The stomach, in the lab. |
| Sulphuric acid | H_2SO_4 | Strong | Acid rain, car batteries, the lab. |
| Nitric acid | HNO_3 | Strong | Acid rain, in the lab. |
| Ethanoic (acetic) acid | CH_3COOH | Weak | Vinegar |
| Methanoic (formic) acid | HCOOH | Weak | Ant & nettle stings, descalers |
| Citric Acid | $\text{C}_6\text{H}_8\text{O}_7$ | Weak | Citrus fruits |

- Acids taste sour (e.g. vinegar, lemon juice).
- Acids are harmful to living cells.
- Aqueous solutions of all acids contain **hydrogen ions, H^+** .

Acids Bases And Salts Study Guide

H. Stephen Stoker, Danny V. White



Acids Bases And Salts Study Guide:

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Study Guide with Answers to Selected Problems

H. Stephen Stoker, Danny V. White, 2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises

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Henold,1984 **A Study Guide to Chemical Principles** Wilbert Hutton,1970 *Study Guide for General Chemistry for Colleges* Herman Thompson Briscoe,1936 Fundamentals of Medicinal Chemistry and Drug Metabolism M. O. Faruk Khan,Ashok Philip,2018-06-01 The primary objective of this 4 volume book series is to educate PharmD students on the subject of medicinal chemistry The book set serves as a reference guide to pharmacists on aspects of chemical basis of drug action This first volume of the series is comprised of 8 chapters focusing on basic background information about medicinal chemistry It takes a succinct and conceptual approach to introducing important fundamental concepts required for a clear understanding of various facets of pharmacotherapeutic agents drug metabolism and important biosynthetic pathways that are relevant to drug action Notable topics covered in this first volume include the scope and importance of medicinal chemistry in pharmacy education a comprehensive discussion of the organic functional groups present in drugs and information about four major types of biomolecules proteins carbohydrates lipids nucleic acids and key heterocyclic ring systems The concepts of acid base chemistry and salt formation and their applications to the drug action and design follow thereafter These include concepts of solubility and lipid water partition coefficient LWPC isosterism stereochemical properties mechanisms of drug action drug receptor interactions critical for pharmacological responses of drugs and much more Students and teachers will be able to integrate the knowledge presented in the book and apply medicinal chemistry concepts to understand the pharmacodynamics and pharmacokinetics of therapeutic agents in the body

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