

# Acid Dissociation Constant (Ka)

Acid Name	Acid Formula	Acid Dissociation Constant (Ka)	pKa
Perchloric acid	$\text{HClO}_4$	$1.6 \times 10^{-11}$	-15.2
Hydroiodic acid	HI	$3.2 \times 10^{-9}$	-9.5
Hydrobromic acid	HBr	$1 \times 10^{-9}$	-9
Hydrochloric acid	Cl	$1.0 \times 10^{-7}$	-6.11
Sulfuric acid	$\text{H}_2\text{SO}_4$	$1.0 \times 10^{-2}$	-3
Hydronium	$\text{H}_3\text{O}^+$	1	0
Citric acid	$\text{C}_6\text{H}_5\text{O}_7$	$7.4 \times 10^{-5}$	3.13
Hydrofluoric acid	HF	$6.6 \times 10^{-4}$	3.18
Formic acid	HCOOH	$1.8 \times 10^{-4}$	3.74
Benzoic acid	$\text{C}_6\text{H}_5\text{COOH}$	$6.3 \times 10^{-5}$	4.20
Acetic acid	$\text{CH}_3\text{COOH}$	$1.8 \times 10^{-5}$	4.74
Carbonic acid	$\text{H}_2\text{CO}_3$	$4.3 \times 10^{-7}$	6.37
Hydrogen sulfide	$\text{H}_2\text{S}$	$9.5 \times 10^{-8}$	7.02
Hypochlorous acid	HOCl	$3.5 \times 10^{-8}$	7.46
Phenol	$\text{C}_6\text{H}_5\text{OH}$	$1.0 \times 10^{-10}$	10

Acids dissociate more  $\Rightarrow$  Stronger acids



$$K_a = \frac{[\text{H}_3\text{O}^+] [\text{A}^-]}{[\text{HA}]}$$

$$pK_a = - \log_{10}(K_a)$$

$[\text{HA}]$  : Undissociated acid concentration

$[\text{H}_3\text{O}^+]$  : Hydronium ion concentration

$[\text{A}^-]$  : Conjugate base concentration

Acids dissociate less  $\Rightarrow$  Weaker acids

## Acid 4 198

**Cavendish Society (London)**

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**Stability Constants of Metal-ion Complexes** International Union of Pure and Applied Chemistry. Commission on Equilibrium Data,1979 *A Comprehensive Treatise on Inorganic and Theoretical Chemistry* Joseph William Mellor,1929 *The Chemical Trade Journal and Chemical Engineer ,1908* **Watts' Dictionary of Chemistry** Henry Watts, Henry Forster Morley,Matthew Moncrieff Pattison Muir,1889 **Drug Metabolism Reviews ,1985** **Proceedings of the American Pharmaceutical Association ,1930** *Index to Gmelin's Handbook of Chemistry ,1872* *The Coal-tar-colours ,1896* **Hand-book of Chemistry** Cavendish Society (London),1872 **Works of the Cavendish Society: Gmelin, Leopold. Hand-book of chemistry. 18 v. & index. 1848-72 ,1872**

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