

Acid Name	Acid Formula	Acid Dissociation Constant (Ka)	pKa
Perchloric acid	HClO <sub>4</sub>	1.6 x 10 <sup>13</sup>	-15.2
Hydroiodic acid	HI	3.2 x 10 <sup>9</sup>	-9.5
Hydrobromic acid	HBr	1 x 10 <sup>9</sup>	-9
Hydrochloric acid	Cl	1.0 x 10 <sup>6</sup>	-6.11
Sulfuric acid	H <sub>2</sub> SO <sub>4</sub>	1.0 x 10 <sup>3</sup>	-3
Hydronium	H <sub>3</sub> O <sup>+</sup>	1	0
Citric acid	C <sub>6</sub> H <sub>8</sub> O <sub>7</sub>	7.4 x 10 <sup>-4</sup>	3.13
Hydrofluoric acid	HF	6.6 x 10 <sup>-4</sup>	3.18
Formic acid	HCOOH	1.8 x 10 <sup>-4</sup>	3.74
Benzoic acid	C <sub>6</sub> H <sub>5</sub> COOH	6.3 x 10 <sup>-5</sup>	4.20
Acetic acid	CH <sub>3</sub> COOH	1.8 x 10 <sup>-5</sup>	4.74
Carbonic acid	H <sub>2</sub> CO <sub>3</sub>	4.3 x 10 <sup>-7</sup>	6.37
Hydrogen sulfide	H <sub>2</sub> S	9.5 x 10 <sup>-8</sup>	7.02
Hypochlorous acid	HOCl	3.5 x 10 <sup>-8</sup>	7.46
Phenol	C <sub>6</sub> H <sub>5</sub> OH	1.0 x 10 <sup>-10</sup>	10



Acids dissociate more => Stronger acids



$$K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$$

$$\text{p}K_a = -\log_{10}(K_a)$$

[HA] : Undissociated acid concentration

[H<sub>3</sub>O<sup>+</sup>] : Hydronium ion concentration

[A<sup>-</sup>] : Conjugate base concentration

Acids dissociate less => Weaker acids

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