



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Explain how the concentration of hydrogen ions in a solution is measured.
- Describe strong and weak acids.

Key Vocabulary

<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base	<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base	<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base
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Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

Key Figures

- Figure 1: pH scale
- Figure 2: Ionization of water

Principles of Ionization

Water is a neutral molecule in which the hydrogen and oxygen atoms are covalently bonded together. It is the ionization of these atoms that gives water its acidic and basic properties.

Water ionizes to form a hydrogen ion and a hydroxide ion. The concentration of these ions is 10^{-7} M at 25°C . The concentration of hydrogen ions is 10^{-7} M at 25°C .

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Acid-Base Properties

<ul style="list-style-type: none">strong acidweak acidstrong baseweak baseacidbaseconjugate acidconjugate base	<ul style="list-style-type: none">strong acidweak acidstrong baseweak baseacidbaseconjugate acidconjugate base	<ul style="list-style-type: none">strong acidweak acidstrong baseweak baseacidbaseconjugate acidconjugate base
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Maurice Bleifeld

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