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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8 3 mole h<sub>2</sub>s and multiply it with 2 mole h<sub>2</sub>o over 2 mole h<sub>2</sub>s from the equation  

$$8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$$
  
 8.3 mole h<sub>2</sub>s x 18g h<sub>2</sub>o i mole h<sub>2</sub>o 149 4g h<sub>2</sub>o is your theoretical yield so percent yield 137 1g h<sub>2</sub>o actual yield 149 4g h<sub>2</sub>o x 100  
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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$
  
 if the reaction of 21 3 grams of o<sub>2</sub> produces a 47 3 yield how many grams of co would be produced step one figure out how many moles of the limiting reagent you have  

$$\frac{21.3 \text{ grams O}_2}{32 \text{ grams mole}} = 0.66 \text{ moles O}_2$$

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$$\text{calculated mass of the product} \times 100$$
  
 in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction  

$$\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$$
  
 4 al 3 o<sub>2</sub> 2 al<sub>2</sub>o<sub>3</sub> b determine the theoretical yield of al<sub>2</sub>o<sub>3</sub> 160 0

g al x 1 mol al x 2 mol al<sub>2</sub>O<sub>3</sub> x 101 964 g al<sub>2</sub>O<sub>3</sub> 302 3 g al<sub>2</sub>O<sub>3</sub> 26 982 g al 4 mol al 1 mol al<sub>2</sub>O<sub>3</sub> c determine the percent yield  
 yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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 weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12  
 c is the limiting reagent 4 33 g of h 2 are left over 13

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 from a set of reactants calculate how much product will be produced from the limiting reagent calculate how much reactant s  
 remains when the reaction is complete

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web the percent yield is calculated as follows text percent yield  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent  
 yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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 how much product will be produced from the limiting reactant calculate how much reactant s remains when the reaction is  
 complete

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 ch<sub>3</sub>oh can be produced by the following reaction 2h<sub>2</sub> co ch<sub>3</sub>oh calculate the theoretical yield of ch<sub>3</sub>oh if 68 5 g of co is  
 reacted with 8 6 g of h<sub>2</sub>

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 reaction

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 which reactant is limiting assuming we started with 30 0 grams of ammonium nitrate and 50 0 grams of sodium phosphate  
 what is the

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80 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2$   $\times$  1 333.8 g  $\text{I}_2$  1 mol  $\text{I}_2$  28 g  $\text{CO}$  1 mol  $\text{CO}$

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