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 $\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu(NO}_3)_2$

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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8.3 \text{ mole H}_2\text{S} \times 2 \text{ mole H}_2\text{O} / 2 \text{ mole H}_2\text{S} = 16.6 \text{ mole H}_2\text{O}$  is your theoretical yield so percent yield  $16.6 \text{ g H}_2\text{O} / 14.0 \text{ g H}_2\text{O} \times 100 = 91.77\%$  problems 6-8 are similar to problem 5

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web limiting reagents problem one 1 for the balanced equation  $\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$  if the reaction of 21.3 grams of  $\text{O}_2$  produces a 47.3% yield how many grams of  $\text{CO}_2$  would be produced step one figure out how many moles of the limiting reagent you have  $0.216 \text{ mol O}_2$  32 grams mole 21.3 grams 0.66 moles 32 g/mol

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web a write a balanced equation for the reaction  $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0

g al x 1 mol al x 2 mol al<sub>2</sub>O<sub>3</sub> x 101 964 g al<sub>2</sub>O<sub>3</sub> 302 3 g al<sub>2</sub>O<sub>3</sub> 26 982 g al 4 mol al 1 mol al<sub>2</sub>O<sub>3</sub> c determine the percent yield yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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$80.0 \text{ g } i_2o_5 \times 1 \text{ mol } i_2o_5 \times 1 \text{ mol } i_2 = 333.8 \text{ g } i_2$

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$\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$

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