

19.5

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of a given acid or base solution.
2. Calculate pH values and concentrations of acids and bases and explain the relationship between pH and pOH.
3. Calculate the pH of a buffer (see problems).

Vocabulary

- | | | |
|-------------------|-----------------------------|--------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. buffer solution |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the requested unit or numerical information of the values indicated that are indicated in the partial equations below. Complete each item completely as indicated.

1. What substance is _____, a strong hydrogen ion donor?
2. _____ is the concentration of H^+ ions present in a solution and equal to _____.
3. _____ is the concentration of _____ in a solution.
4. _____ is the concentration of _____ in a solution.
5. _____ is the concentration of _____ in a solution.
6. _____ is the concentration of _____ in a solution.
7. _____ is the concentration of _____ in a solution.
8. _____ is the concentration of _____ in a solution.
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16. _____ is the concentration of _____ in a solution.
17. _____ is the concentration of _____ in a solution.
18. _____ is the concentration of _____ in a solution.
19. _____ is the concentration of _____ in a solution.
20. _____ is the concentration of _____ in a solution.

1. _____
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20. _____

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