

12: Stoichiometry

Key Stoichiometry Terms

- Stoichiometry:** Using the mole ratio in the balanced equation and information about one compound to find information about another in the reaction.
- Dimensional Analysis:** Method of converting units by multiplying by ratio of equalities.
- Molar Mass:** Sum of all the atomic masses (from the periodic table) in the compound.
- Solute:** Substance being dissolved in a homogeneous mixture (solution).
- Concentration:** Measure of how much solute is dissolved in how much solution.
- Molarity (M):** A concentration unit in moles per unit liter.
- Standard Temperature and Pressure (STP):** 1 atm (or 101.3 kPa) and 273 K (0°C).
- Molar volume of a gas:** 1 mole of any gas at STP is 22.4 liters.
- Limiting Reactant:** Reactant that stops the reaction by running out first.
- Actual Yield:** The amount actually produced in the lab.
- Theoretical Yield:** Amount that should be theoretically produced based on stoichiometric calculations.
- Percent Yield:** Compares the actual yield to the theoretical yield (their ratio with %).

Chemical Equations

- Coefficients give the mole ratios.

Example:



2 moles of H₂ react with 1 mole of O₂ and produces 2 moles of H₂O.

Dimensional Analysis in Stoichiometry

Equalities used during dimensional analysis for stoichiometry:

- Mole ratio in balanced equation:** Use to convert between moles of different compounds in the balanced equation.
- Molar mass:** Used to convert between grams and moles.
- Concentration:** Used to convert between moles and liters of a solution.
- Molar volume of a gas:** Used to convert between moles and liters of a gas at STP.

KUDOS for Stoichiometric Problems

Method to solve word problems:

- K = Known:** Identify all the known information (given quantities, balanced equation, etc.).
- U = Unknown:** What quantity and what compound does the problem ask for?
- D = Definition:** Identify the equalities that will be needed during dimensional analysis to convert from the known to the unknown.
- O = Output:** Perform the dimensional analysis.
- S = Substantiation:** Check for reasonableness, units and correct significant figures.

Molarity

Molarity - the concentration measure by the moles of solute per liter of solution.

$$\text{Molarity} = \frac{\text{moles solute}}{\text{L solution}}$$

Limiting Reactants

- Once a reactant has run out, the reaction will stop.
- Do stoichiometry for each given reactant quantity to the same product each time. Choose the calculation that gives the **smallest amount** of product.
- The reactant that produced the **smallest amount** of product is the **limiting reactant**.
- Steps to determine the limiting reactant: grams reactant \rightarrow moles reactant \rightarrow moles product \rightarrow grams product. Alternatively, use the molar method below.

Simple Limiting Reagent Finder Mnemonic: Find the moles of each reactant and divide the moles of each reactant by its coefficient. The reactant with the **smallest number** is the limiting reagent = "Smallest m/c is limited".

Percent Yield

$$\% \text{ yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

Example:

If stoichiometry calculations tell you that a reaction should produce 10.5 g, but when you perform the lab you produce 8.7 g, what was the percent yield?

$$\% \text{ yield} = \frac{8.7 \text{ g}}{10.5 \text{ g}} \times 100 = 82.9\%$$

Stoichiometry Examples

All examples for the following reaction:



Mole-Mass examples:

0.75 moles HCl react.	How many grams Mg are needed?
0.75 moles HCl	1 mole Mg $\frac{24.31 \text{ g Mg}}{2 \text{ mole HCl}}$ $= 12.15 \text{ g Mg}$

Mass-Mass example:

17.25 g Mg react.	How many grams MgCl ₂ are produced?
17.25 g Mg	1 mole Mg $\frac{95.21 \text{ g MgCl}_2}{1 \text{ mole Mg}}$ $= 95.21 \text{ g MgCl}_2$

Mass-Liter example (solution):

How many liters of 1.7 M HCl are needed to react with 2.5 g Mg?	2.5 g Mg	1 mole Mg	2 mole HCl	1 L HCl
	2.5 g Mg	1 mole Mg	$\frac{1.7 \text{ mole HCl}}{1 \text{ mole Mg}}$	$= 0.12 \text{ L HCl}$

Mass-Liter example (gas):

How many liters of H ₂ are produced when 3.5 g Mg react?	3.5 g Mg	1 mole Mg	1 mole H ₂	22.4 L H ₂
	3.5 g Mg	1 mole Mg	$\frac{1 \text{ mole H}_2}{1 \text{ mole Mg}}$	$= 3.5 \text{ L H}_2$

Limiting reactant:

If 0.25 mole HCl reacts with 0.55 mole Mg, how many grams of MgCl₂ are produced and what was the limiting reactant?

2HCl + Mg \rightarrow MgCl ₂ + H ₂	0.25 mole HCl	1 mole HCl	$\frac{95.21 \text{ g MgCl}_2}{1 \text{ mole HCl}}$	95.21 g MgCl ₂
	0.25 mole HCl	2 mole HCl	$\frac{1 \text{ mole MgCl}_2}{2 \text{ mole HCl}}$	(none)

0.55 mole Mg	1 mole Mg	$\frac{95.21 \text{ g MgCl}_2}{1 \text{ mole Mg}}$	95.21 g MgCl ₂
	0.55 mole Mg	$\frac{1 \text{ mole MgCl}_2}{1 \text{ mole Mg}}$	$= 52.6 \text{ g MgCl}_2$

1.2 g MgCl₂ will be produced and HCl is the limiting reactant.

How to Use This Cheat Sheet: These are the keys related to this topic. Try to read through it carefully twice then write it out from memory on a blank sheet of paper. Review it again before the exam.

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I. Dwaine Eubanks, Lucy T. Eubanks

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Preparing for Your ACS Examination in Organic Chemistry

I. Dwaine Eubanks, Lucy T. Eubanks, 2002-01-01

Preparing for Your ACS Examination in Organic Chemistry Examinations Institute-American Chemical Society Division of Chemical Education, 2019-12

Organic Chemistry Study Guide Subject Guide to Books in Print, 2001 **Making it relevant** Peter Nentwig, David Waddington, 2006

Teaching in context has become an accepted and often welcomed way of teaching science in both primary and secondary schools. The conference organised by IPN and the University of York Science Education Group Context based science curricula drew on the experience of over 40 science educators and 10 projects. The book is arranged in four parts. Part A consists of two papers one on situated learning and the other on implementation of new curricula. Part B contains descriptions of five major curricula in different countries why they were introduced, how they were developed and implemented and evaluation results. Part C gives descriptions of three projects that are of smaller scale and their materials are used as interventions in other more conventional curricula. There is also a contribution on some fundamental research where modules of work are written to examine how best to design context based curricula. Finally Part D consists of two chapters one summarising some of the findings that came out of the chapters in the three earlier parts and the second looks at the future.

Whole-class Inquiry Dennis W. Smith, Henry Joan Gallager-Bolos, 2009

In response to requests from science education professionals this is the perfect vehicle for implementing and assessing this concept of whole class inquiry in your classroom. This is a must have package for preservice and inservice middle and high school science teachers.

Preparing for Your ACS Examination in Physical Chemistry Thomas A. Holme, Kristen Murphy, 2009

The American Chemical Society at 125, 2002

10 Things You Must Know About The International Chemistry Olympiad (Icho): A Guide To The Icho Competition (Revised Edition) I-jy Chang, Fun Man Fung, 2023-02-17

This is a book dedicated to special groups in the community whom we reach out to promote chemistry interest and learning chemistry. It will be immensely useful to students.

parents who are interested in Science Olympiads Chemistry Olympiad chemistry instructors at the secondary and collegiate level and other science instructors at the secondary and collegiate level This is a one of a kind book that details the behind the scenes preparation and trainings that students undergo in various countries It will enable the effective promotion of chemistry in your respective countries and around the globe Preparing for Your ACS Examination in General Chemistry - the Official Guide LUCY T. EUBANKS,2018-10-15 This guide is separated into first term and second term general chemistry material Each section contains 8 chapters of material that also aligns to most general chemistry textbooks for a seamless addition to study materials for students Each chapter is designed with an introductory section of the material including common representations and where to find this material in a textbook The second section provides worked examples of typical multiple choice questions including how the correct answer is determined as well as how the incorrect answers were determined Also included for each study problem is a listing of the corresponding practice questions that use that concept The final section is a series of practice problems to test the concepts collectively The key is provided on a separate page for all study and practice problems Science Tests and Reviews Buros Center,1975 Science Tests and Reviews consisting of science sections of the first seven MMYs and Tests in Print II includes 217 original test reviews written by 81 specialists 18 excerpted test reviews 270 references on the construction use and validity of specific tests a bibliography on in print science tests references for specific tests cumulative name indexes for specific tests with references a publishers directory title index name index and a scanning index The 97 tests covered fall into the following categories 23 general 14 biology 35 chemistry 3 geology 6 miscellaneous and 16 physics POGIL Shawn R. Simonson,2023-07-03 Process Oriented Guided Inquiry

Learning POGIL is a pedagogy that is based on research on how people learn and has been shown to lead to better student outcomes in many contexts and in a variety of academic disciplines Beyond facilitating students mastery of a discipline it promotes vital educational outcomes such as communication skills and critical thinking Its active international community of practitioners provides accessible educational development and support for anyone developing related courses Having started as a process developed by a group of chemistry professors focused on helping their students better grasp the concepts of general chemistry The POGIL Project has grown into a dynamic organization of committed instructors who help each other transform classrooms and improve student success develop curricular materials to assist this process conduct research expanding what is known about learning and teaching and provide professional development and collegiality from elementary teachers to college professors As a pedagogy it has been shown to be effective in a variety of content areas and at different educational levels This is an introduction to the process and the community Every POGIL classroom is different and is a reflection of the uniqueness of the particular context the institution department physical space student body and instructor but follows a common structure in which students work cooperatively in self managed small groups of three or four The group work is focused on activities that are carefully designed and scaffolded to enable students to develop important

concepts or to deepen and refine their understanding of those ideas or concepts for themselves based entirely on data provided in class not on prior reading of the textbook or other introduction to the topic. The learning environment is structured to support the development of process skills such as teamwork, effective communication, information processing, problem solving and critical thinking. The instructor's role is to facilitate the development of student concepts and process skills, not to simply deliver content to the students. The first part of this book introduces the theoretical and philosophical foundations of POGIL pedagogy and summarizes the literature demonstrating its efficacy. The second part of the book focusses on implementing POGIL, covering the formation and effective management of student teams, offering guidance on the selection and writing of POGIL activities as well as on facilitation, teaching large classes and assessment. The book concludes with examples of implementation in STEM and non STEM disciplines as well as guidance on how to get started. Appendices provide additional resources and information about The POGIL Project.

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