

**19****ACIDS, BASES, AND SALTS****SECTION 19.1 ACID-BASE THEORIES (pages 587–593)**

*This section compares and contrasts acids and bases as defined by the theories of Arrhenius, Brønsted-Lowry, and Lewis. It also identifies conjugate acid-base pairs in acid-base reactions.*

**► Properties of Acids and Bases (pages 587–588)**

1. Circle the letters of all the terms that complete the sentence correctly.  
The properties of acids include \_\_\_\_\_.
  - a. reacting with metals to produce oxygen
  - b. giving foods a sour taste
  - c. forming solutions that conduct electricity
  - d. causing indicators to change color
2. Bases are compounds that react with acids to form \_\_\_\_\_ and a(n) \_\_\_\_\_.
3. Circle the letters of all the terms that complete the sentence correctly.  
The properties of bases include \_\_\_\_\_.
  - a. tasting bitter
  - b. feeling slippery
  - c. changing the color of an indicator
  - d. always acting as a strong electrolyte

**► Arrhenius Acids and Bases (pages 588–590)**

4. Match the number of ionizable hydrogens with the type of acid.

_____ one	a. diprotic
_____ two	b. triprotic
_____ three	c. monoprotic
5. Is the following sentence true or false? Only the hydrogens in weak polar bonds are ionizable. \_\_\_\_\_
6. Hydrogen is joined to a very \_\_\_\_\_ element in a very polar bond.
7. Alkali metals react with water to produce \_\_\_\_\_ solutions.

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**U.S. Atomic Energy Commission**

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