

## ACIDS, BASES, & THE pH SCALE

Name: Answer Key

Date: \_\_\_\_\_

Pd: \_\_\_\_\_

### PART 1: ACIDS AND BASES

Comparing Acids and Bases	
Acids	Bases (Alkalines)
<ul style="list-style-type: none"><li>• <u>Sour</u> taste, leaves <u>burning</u> sensation</li><li>• <u>Increase</u> hydrogen ions (<math>H^+</math>) in solution</li><li>• Example: <u><math>HCl</math></u>      <u><math>H^+ + Cl^-</math></u></li><li>• Turns <u>litmus</u> (pH) paper to <u>red</u></li></ul>	<ul style="list-style-type: none"><li>• Sharp <u>bitter</u> taste, feels <u>slippery</u> to touch</li><li>• <u>Decrease</u> hydrogen ions (<math>H^+</math>) in solution (<u>Increase</u> hydroxide ions (<math>OH^-</math>) in solution)</li><li>• Example: <u><math>NaOH</math></u>      <u><math>Na^+ + OH^-</math></u></li><li>• Turns litmus (pH) paper to <u>blue</u></li></ul>

### PART 2: THE pH SCALE

- The pH Scale is a measurement of how acidic or how basic a solution is.
- The scale goes from 0 to 14.
- Compounds are acidic if they have a pH lower than 7.
- Compounds with a pH higher than 7 are basic or alkaline.
- Compounds with a pH of 7, are neutral.

0                      1           2           3           4           5           6           7           8           9           10           11           12           13                      14

Increased concentration of  $H^+$  ions  
Decreased concentration of  $OH^-$  ions

Decreased concentration of  $H^+$  ions  
Increased concentration of  $OH^-$  ions

pH below 7 = acids

pH at 7 = neutral  
(pure water)

pH above 7 = bases

### PART 3: BUFFERS

- A buffer is a weak acid or base that prevents sudden changes in pH.
  - Maintains homeostasis
- A buffer works by binding to  $H^+$  ions releasing  $H^+$  ions when the solution pH changes.
- Buffers bind to  $H^+$  ions when the concentration of  $H^+$  ions in solution increases.
- $NaOH$        $Na^+ + OH^-$  (The  $OH^-$  will bind to  $H^+$  ions to form  $H_2O$  - neutralization)
- Buffers release  $H^+$  ions when the concentration of  $H^+$  ions in solution decreases.  
 $HCl$        $H^+ + Cl^-$

# Acids And Bases Study Guide Key

**Karen Timberlake**



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*pH buffering system* Mário Sérgio Piantavini, 2025-09-19 Acid Base Balance A Comprehensive Study Guide explores the fundamental principles of pH regulation and buffer systems in human physiology The book presents a historical overview of acid base research from the early 20th century to modern developments followed by an in depth study of key buffering mechanisms the role of lungs and kidneys in homeostasis and the diagnostic use of blood gas analysis It examines metabolic and respiratory disorders their compensatory responses and the complexities of mixed acid base disturbances Special clinical contexts such as cirrhosis hemodialysis and neonatal physiology are addressed highlighting the challenges of maintaining acid base stability in vulnerable populations The book also introduces modern analytical technologies mathematical models and innovative monitoring systems bridging theory and clinical application This guide serves as a valuable resource for students healthcare professionals and researchers seeking a clear yet thorough understanding of acid base physiology and its clinical implications

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