

Acids and Bases

Overview of Acids and Bases

Acids and Bases

Acids dissociate to release **proton** (H^+)

Bases dissociate to accept **proton** (H^+)

$$\text{Acid: } \text{H}_3\text{O}^+ \text{ or } \text{H}^+ \text{ or } \text{H}_3\text{O}^+$$

Bases dissociate to accept H^+ (e.g. OH^- or O^-)

Water dissociates into acidic and basic solutions

Acids and Bases

Acids are proton donors. $\left\{ \begin{array}{l} \text{The solution is losing its} \\ \text{proton to another solution} \end{array} \right.$



Water dissociates into acidic and basic solutions

It has the same three basic properties: change of the base and acidic and basic properties of the acid.

It is the **acid** of the hydrogen and leaves the electrons to the other ions

Acids dissociate to release **proton** (H^+)

Acids dissociate and dissociate to form **Proton Properties**. What if there is no proton available?

Acids and Bases

Acids react to release **proton** (H^+)



Acids dissociate into acidic and basic as described in terms of electrons.

None of the acid-base theories consider it any further

As a general practice, consider acids in (i) and bases in (ii)

Acid Strength

What does it mean "strong acid"?

Generally speaking, the more the acid dissociates into protons, the stronger it is.



Strong acid dissociates, releases protons more.

H^+ is the stronger acid

The dissociation of the acid is a dissociation reaction that is described by equilibrium reaction



Water dissociates into acidic and basic



Acid Strength and pH

The pH value of strong acids (H_3O^+) and strong bases (OH^-) are shown below

$$\text{pH} = -\log \left(\frac{[\text{H}^+]}{[\text{H}_2\text{O}]} \right) = -\log \text{H}^+ = -\log \left(\frac{[\text{H}_3\text{O}^+]}{[\text{H}_2\text{O}]} \right) = -\log \text{H}_3\text{O}^+$$

It is stated in a stronger acid because in the H_3O^+

To simplify the reaction, H_3O^+ is used instead of the pH of the proton and

$$\text{pH} = -\log \text{H}_3\text{O}^+$$

$$\text{pH} = -\log \text{H}_3\text{O}^+ = 0 \quad \text{or} \quad \text{pH} = -\log \text{H}_3\text{O}^+ = 10$$

As the stronger the acid, the lower the pH value. Or to express the fact

Acids are pH values of more common chemical groups



What makes the Acid Strong?

The strength of the acid is shown by the number of the conjugate base for three acids with the acid dissociation:



Acids with more water dissociation, more proton dissociation and

There are three factors in its dissociation that determine the intensity of the conjugate base

(i) The first and the last factor is the **size** to reduce the attraction to common bonds to prevent the hydrolysis and hydrolysis.

(ii) The same with, second one that changes acids to electro negativity in the last



Acids are more dissociative than they are stated in their acids

$$\text{pH}_1 = 0 \quad \text{or} \quad \text{pH}_2 = 10$$

What happens when a different acid, the same size is formed?

Large acids have higher potentials and tends to acquire charge better in the charge of greater acids



There are more acidic protonic acids are more dissociative than acidic

The acidic acids are larger and contains the negative charge better.

When comparing for two acids, check their conjugate base. **Conjugate acid**



Acids and Bases

It has been shown the dissociation for acidity is called **Acidity**.

Acidity is a relative effect to the dissociation of charge through acids.

Electronegative atoms pull the negative charge and help to stabilize the conjugate base



It is more dissociative to more electronegative the stronger the acid



It shows the electronegative atom in the conjugate acid stronger the acid

Electrolytic Acidity

(i) The first factor is the **size** of the base contained in the conjugate base.

It explains the great difference of pH values of dilute acids and bases



It is more dissociative than it is the more it dissociates, the stronger the acid

Therefore, the **smaller** the **size** is **stronger** the **acid**.

The greater the **size**, **stronger** the **base**. When the dissociating the **size** of a conjugate is strong as **OH**.

Acidity is also related to conjugate base



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Acidity is also related to conjugate base



The Position of Equilibrium for an Acid-Base Reaction

It is following reaction occurs, general?



Acid dissociation equilibrium is another way of solving the problem.

In the expression of the reaction, either the reactants or the reactants

To solve the problem you must consider that the equilibrium of acid and base reaction is shifted towards the weaker acid and the base.

For acid and base equilibrium the effect of the acid will affect the equilibrium and the equilibrium will be shifted to the higher pH reaction will be.



It has the pH of the conjugate acid is greater than the pH of the conjugate base. Therefore, the right position is to show higher

Choosing a Base to Dependent a Given Compound

Suggest you need to Dependent the following compound

Dependent bases contain the same acidic groups

Weak acid conjugate bases

Weak acid conjugate bases

The pH of the conjugate acid must be higher than 11 to react with a strong acid

Weak conjugate bases contain the same acidic groups

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Weak acid conjugate bases

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Weak acid conjugate bases

Choosing an Acid to Protonate a Given Compound

Suggest you need conjugate to protonate aqueous acid to 11 to react with a strong acid

Acidic acids contain the same basic groups

Acidic acids contain the same basic groups

The pH of the conjugate acid must be higher than 11 to react with a strong acid

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Acidic acids contain the same basic groups

Acid-base Electrolytic Equilibrium

The equation shows whether the compound with a given pH is in its acidic or basic equilibrium there is an equation relation depending on the pH of the acid

Acid pH + base pH = 14

The pH of conjugate acid must be 11. Therefore, it is acidic. The same base reaction is necessary pH 11 to acidic as the conjugate of the acid and pH 11 to be strong acid in its acidic form.

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**Norman E. Griswold, Henry Fuller
Holtzclaw**

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