

**Guided Practice - Acids & Bases****Properties of Acids and Bases****1. Match the following terms with their descriptions:**

a. Autoionized Acid \_\_\_\_\_ the substance formed when a proton is added to \_\_\_\_\_ a base

b. Autoionized Base \_\_\_\_\_ the remaining substance when a proton is taken from an acid

c. Conjugate Acid \_\_\_\_\_ a substance that produces  $\text{OH}^-$  in solution

d. Conjugate Base \_\_\_\_\_ a substance producing  $\text{H}^+$  in solution

e. Bronsted-Lowry Acid \_\_\_\_\_ a proton acceptor

f. Bronsted-Lowry Base \_\_\_\_\_ a proton donor

**2. Which of the following represent conjugate acid-base pairs? Label each species as a conjugate acid or a conjugate base. For those pairs that are NOT conjugates, write the correct conjugate acid or base for each species in the pair, labeling each species correctly.**

a.  $\text{HSO}_4^-$ ,  $\text{SO}_4^{2-}$

b.  $\text{HBr}$ ,  $\text{BrO}^-$

c.  $\text{H}_2\text{PO}_4^-$ ,  $\text{PO}_4^{3-}$

d.  $\text{H}_2\text{O}^{2-}$ ,  $\text{HO}^{2-}$

**3. Complete the following table regarding acid strength.**

Property/Condition	Strong Acid	Weak Acid
Conjugate base		
Species present in solution		
Relative strengths of conjugate base compared to water		
Equations following acids in order from strongest		

# Acids And Bases Answer Packet

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Farrell**

## Acids And Bases Answer Packet:

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