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# **Inorganic Chemistry ACS: Study Guide Exam|334 Questions with Accurate Answers**

Ionization energy - -energy required to remove the least tightly bound electron from a neutral atom in the gas phase

- periodic trend of ionization energy - -highest at top right-smaller electron=harder to remove

- Why is a half filled subshell so stable? - -it serves to maximize the stabilizing interactions while minimizing the destabilizing interactions among electrons

- exchange interaction - -pie, stabilizing, result of electrons pairing in degenerate orbitals with parallel spin

- pairing energy - -destabilizing, coulomb interaction, pic, energy of electron-electron repulsion in a filled orbital

- Is it easier to ionize a high energy or low energy electrons - -high energy electron-already contains more energy so it requires less energy input

- What happens when a 3d series metal is ionized? - -the first electron to be ionized will come from the 4s orbital, the other s electron will enter the d orbital ( $4s^0 3d^{n+1}$ )

- lanthanide contraction - -reduction in atomic radius following the lanthanide series, contrary to the overall trend observed for the periodic table

- lanthanides - -elements 57-71, first appearance of f orbitals, f orbitals are poor at shielding so any electrons added will have a higher  $Z_{eff}$ , shrinking the radius

- Slater's rules - -tell us what the effective nuclear charge will be,  $Z_{eff} = Z - \sigma$ , Z is the atomic number,  $\sigma$ =sum of the number of electrons in a given subshell multiplied by a weighting coefficient (page 1)

- Shielding - -the reduction in charge attraction between the nucleus and electrons due to electrons between the nucleus and the electron in question, it is considered the be between if it has a lower energy

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